

Balancing Equations Worksheet

- 1) $\text{___ C}_3\text{H}_8 + \text{___ O}_2 \rightarrow \text{___ CO}_2 + \text{___ H}_2\text{O}$
- 2) $\text{___ Al} + \text{___ Fe}_3\text{N}_2 \rightarrow \text{___ AlN} + \text{___ Fe}$
- 3) $\text{___ Na} + \text{___ Cl}_2 \rightarrow \text{___ NaCl}$
- 4) $\text{___ H}_2\text{O}_2 \rightarrow \text{___ H}_2\text{O} + \text{___ O}_2$
- 5) $\text{___ C}_6\text{H}_{12}\text{O}_6 + \text{___ O}_2 \rightarrow \text{___ H}_2\text{O} + \text{___ CO}_2$
- 6) $\text{___ H}_2\text{O} + \text{___ CO}_2 \rightarrow \text{___ C}_7\text{H}_8 + \text{___ O}_2$
- 7) $\text{___ NaClO}_3 \rightarrow \text{___ NaCl} + \text{___ O}_2$
- 8) $\text{___ (NH}_4\text{)}_3\text{PO}_4 + \text{___ Pb(NO}_3\text{)}_4 \rightarrow \text{___ Pb}_3\text{(PO}_4\text{)}_4 + \text{___ NH}_4\text{NO}_3$
- 9) $\text{___ BF}_3 + \text{___ Li}_2\text{SO}_3 \rightarrow \text{___ B}_2\text{(SO}_3\text{)}_3 + \text{___ LiF}$
- 10) $\text{___ C}_7\text{H}_{17} + \text{___ O}_2 \rightarrow \text{___ CO}_2 + \text{___ H}_2\text{O}$
- 11) $\text{___ CaCO}_3 + \text{___ H}_3\text{PO}_4 \rightarrow \text{___ Ca}_3\text{(PO}_4\text{)}_2 + \text{___ H}_2\text{CO}_3$
- 12) $\text{___ Ag}_2\text{S} \rightarrow \text{___ Ag} + \text{___ S}_8$
- 13) $\text{___ KBr} + \text{___ Fe(OH)}_3 \rightarrow \text{___ KOH} + \text{___ FeBr}_3$
- 14) $\text{___ KNO}_3 + \text{___ H}_2\text{CO}_3 \rightarrow \text{___ K}_2\text{CO}_3 + \text{___ HNO}_3$
- 15) $\text{___ Pb(OH)}_4 + \text{___ Cu}_2\text{O} \rightarrow \text{___ PbO}_2 + \text{___ CuOH}$
- 16) $\text{___ Cr(NO}_2\text{)}_2 + \text{___ (NH}_4\text{)}_2\text{SO}_4 \rightarrow \text{___ CrSO}_4 + \text{___ NH}_4\text{NO}_2$
- 17) $\text{___ KOH} + \text{___ Co}_3\text{(PO}_4\text{)}_2 \rightarrow \text{___ K}_3\text{PO}_4 + \text{___ Co(OH)}_2$
- 18) $\text{___ Sn(NO}_2\text{)}_4 + \text{___ Pt}_3\text{N}_4 \rightarrow \text{___ Sn}_3\text{N}_4 + \text{___ Pt(NO}_2\text{)}_4$
- 19) $\text{___ B}_2\text{Br}_6 + \text{___ HNO}_3 \rightarrow \text{___ B(NO}_3\text{)}_3 + \text{___ HBr}$
- 20) $\text{___ ZnS} + \text{___ AlP} \rightarrow \text{___ Zn}_3\text{P}_2 + \text{___ Al}_2\text{S}_3$

Balancing Chemical Equations Section 11 1 Answers

John T. Moore, Ed.D.



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Metals and Chemical Change D A Johnson, 2007-10-31 This book looks at how molecules react and how the feasibility and outcome of chemical reactions can be predicted Beginning with an introduction to the concept of an activity series of metals Metals and Chemical Change then introduces chemical thermodynamics enthalpy entropy and free energy and applies the concept to both inorganic and organic elements A Case Study on batteries and fuel cells is also included The accompanying CD ROM includes video sequences of the reactions of metals with water acid and aqueous ions and gives the reader an opportunity to make experimental observations and predictions about chemical behaviour A comprehensive Data Book of chemical and physical constants is included along with a set of interactive self assessment questions The Molecular World series provides an integrated introduction to all branches of chemistry for both students wishing to specialise and those wishing to gain a broad understanding of chemistry and its relevance to the everyday world and to other areas of science The books with their Case Studies and accompanying multi media interactive CD ROMs will also provide valuable resource material for teachers and lecturers The CD ROMs are designed for use on a PC running Windows 95 98 ME or 2000

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