

Chemistry Unit 15: Chemical Equilibrium Worksheet**Practice 15.1 Writing Equilibrium Equations**

Write the equilibrium expressions (K_{eq}) for the following reactions

1. $2\text{CO(g)} \leftrightarrow \text{C(s)} + \text{CO}_2\text{(g)}$
2. $\text{N}_2\text{O}_4\text{(g)} \leftrightarrow 2\text{NO}_2\text{(g)}$
3. $2\text{HCl(g)} \leftrightarrow \text{I}_2\text{(g)} + \text{Cl}_2\text{(g)}$
4. $\text{Fe}_2\text{O}_3\text{(s)} + 3\text{H}_2\text{(g)} \leftrightarrow 2\text{Fe(s)} + 3\text{H}_2\text{O(g)}$
5. $2\text{CO(g)} + 4\text{H}_2\text{O(g)} \leftrightarrow 3\text{O}_2\text{(g)} + 2\text{CH}_4\text{(g)}$
6. $\text{P}_4\text{(s)} + 6\text{NO(g)} \leftrightarrow \text{P}_4\text{O}_6\text{(s)} + 3\text{N}_2\text{(g)}$

15.2 Calculating Equilibrium Constants

Be sure to write out the equilibrium (K_{eq}) expression for each problem before trying to solve it!

1. What is the equilibrium expression for the following reaction?
 $\text{SeO}_2\text{(s)} + 2\text{CO(g)} \leftrightarrow \text{Se(s)} + 2\text{CO}_2\text{(g)}$
2. At equilibrium the following concentrations are recorded: 20.0 M of H_2 , 18.0 M of CO_2 , 12.0 M H_2O , and 5.9 M of CO at 427°C. What is the value of K_{eq} at this temperature for the following reaction?
 $\text{CO}_2\text{(g)} + \text{H}_2\text{(g)} \leftrightarrow \text{CO(g)} + \text{H}_2\text{O(g)}$
3. At equilibrium a mixture of N_2 , H_2 , and NH_3 gas at 500°C is determined to consist of 0.602 M N_2 , 0.420 M H_2 , and 0.113 M of NH_3 . What is the equilibrium constant at this temperature? Does this equilibrium favor the reactants or the products?
 $\text{N}_2\text{(g)} + 3\text{H}_2\text{(g)} \leftrightarrow 2\text{NH}_3\text{(g)}$
4. Given the equilibrium concentrations of 0.15 M NO , 0.12 M H_2 and 1.6 M N_2 . Calculate the equilibrium constant for the following reaction. Does this reaction favor the reactants or the products?
 $2\text{NO(g)} + 2\text{H}_2\text{(g)} \leftrightarrow \text{N}_2\text{(g)} + 2\text{H}_2\text{O(l)}$
5. What is the equilibrium constant for the following reaction if the equilibrium concentrations for the gases are 0.3 M for C_2H_6 , 0.6 M for O_2 , 1.3 M for CO_2 and 1.8 M for H_2O ? Does the reaction favor the reactants or products?
 $2\text{C}_2\text{H}_6\text{(g)} + 7\text{O}_2\text{(g)} \leftrightarrow 4\text{CO}_2\text{(g)} + 6\text{H}_2\text{O(g)}$
6. Nitrogen dioxide reacts to form dinitrogen tetraoxide. What is the equilibrium concentration of N_2O_4 if the $K_{eq} = 4.70$ and the equilibrium concentration of NO_2 is 0.0310 M?
 $2\text{NO}_2\text{(g)} \leftrightarrow \text{N}_2\text{O}_4\text{(g)}$

Calculating Equilibrium Constant Chem Worksheet 18 3

James E. Brady



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Remediation Engineering Suthan S. Suthersan, John Horst, Matthew Schnobrich, Nicklaus Welty, Jeff McDonough, 2016-11-25. This second edition of Remediation Engineering will continue to be the seminal handbook that regulators must have on hand to address any of the remediation issues they are grappling with daily. The book is wide ranging but specific enough to address any environmental remediation challenge. Patricia Reyes, Interstate Technology Regulatory Council, Washington, DC, USA. This book offers the researcher, teacher, practitioner, student, and regulator with state of the art advances in conducting site investigations and remediation for common and emerging contaminants. It is revolutionary in its approach to conducting subsurface investigation which greatly influences a successful and appropriate response in assessing and addressing environmental risk. This book is a giant leap forward in understanding how contaminants behave and how to reduce risk to acceptable levels in the natural world. Daniel T. Rogers, Amsted Industries Incorporated, Chicago, Illinois, USA. This text is a superb reference and a good tool for learning about state of the art techniques in remediation of soil and groundwater. It will become a ready reference at many companies as the engineering community creates increased value from remediation efforts around the world. John Waites, AVX Corporation, Fountain Inn, South Carolina, USA. Remediation Engineering was first published in 1996 and quickly became the go-to reference for a relatively young industry offering the first comprehensive look at the state of the science in treatment technologies of the time and the contaminants they applied to. This fully updated Second Edition will capture the fundamental advancements that have taken place during the last two decades within all the subdisciplines that form the foundation of the remediation engineering platform. It covers the entire spectrum of current technologies that are employed in the industry and also discusses future trends and how practitioners should anticipate and adapt to those needs. Features: Shares the latest paradigms in remediation design approach and contaminant hydrogeology. Presents the landscape of new and emerging contaminants. Details the current state of the practice for both conventional technologies such as sparging and venting. Examines newer technologies such as dynamic groundwater recirculation and injection-based remedies to address both organic and inorganic contaminants. Describes the advances in site characterization concepts such as smart investigations and digital conceptual site models. Includes all new color photographs and figures.

INIS Atomindex, 1984

Geochemistry International, 1998 Vols for 1964 v 2 no 1. 1965 include selected articles translated from geochemical papers from other languages but primarily from Russian, German, French, and Japanese.

Tutorial, Human Factors in Software Development Bill Curtis, 1981. Presents a board overview of the experimental research on human factors in software development.

Introduction *Data Sources*, 2000 *Chemical Equilibrium* William Guenther, 1975-12. The present work is

designed to provide a practical introduction to aqueous equilibrium phenomena for both students and research workers in chemistry biochemistry geochemistry and interdisciplinary environmental fields The pedagogical strategy I have adopted makes heavy use of detailed examples of problem solving from real cases arising both in laboratory research and in the study of systems occurring in nature The procedure starts with mathematically complete equations that will provide valid solutions of equilibrium problems instead of the traditional approach through approximate concentrations and idealized infinite dilution assumptions There is repeated emphasis on the use of corrected conditional equilibrium constants and on the checking of numerical results by substitution in complete equations and or against graphs of species distributions Graphical methods of calculation and display are used extensively because of their value in clarifying equilibria and in leading one quickly to valid numerical approximations The coverage of solution equilibrium phenomena is not however exhaustively comprehensive Rather I have chosen to offer fundamental and rigorous examinations of homogeneous step equilibria and their interactions with solubility and redox equilibria Many examples are worked out in detail to demonstrate the use of equilibrium calculations and diagrams in various fields of investigation

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Chemical Calculations Paul Yates, 1997-06-05 Chemical Calculations provides an introduction to the mathematics required for physical chemistry courses This book is unique in that it provides a gentle introduction with a chemistry centered rather than math centered approach Written by a chemist for undergraduate students it imparts an understanding of the subject from a chemist's viewpoint using examples from real chemistry It includes illustrations that show exactly how to use calculators to work problems and examples of important chemical problems with fully worked solutions This book is an ideal companion throughout a chemistry course that can be consulted when required and used to keep one step ahead of the lecture *Study Guide [to Accompany] General Chemistry* James E. Brady, 1982 Chemical Calculations Sidney William Benson, 1963 Chemical Equilibrium and Solutions James E. Banks, 1967

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