

## Calculating Equilibrium Concentrations

Given the Equilibrium Constant AND the initial conditions, equilibrium concentrations can be calculated.

Example:

The equilibrium constant for the reaction



is 0.021 at a certain temperature. If the initial concentration of  $\text{PCl}_{5(g)}$  is  $0.100 \text{ mol L}^{-1}$ , what are the equilibrium concentrations of all species?

Assume a volume of 1 L.

Initially it contains 0.100 moles  $\text{PCl}_{5(g)}$ .

At equilibrium,  $x$  moles of  $\text{PCl}_{5(g)}$  have reacted and therefore  $0.1-x$  moles are left. From the equation  $x$  moles of  $\text{PCl}_{3(g)}$  and of  $\text{Cl}_{2(g)}$  have been formed.

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**S Ben Porath**



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