Chemistry Unit 15: Chemical Equilibrium Worksheet

Practice 15.1 Writing Equilibrium Equations

Write the equilibrium expressions (K_{oo}) for the following reactions

- 2CO(g) ← C(s) + CO₂(g)
- N₂O_b(g) ↔ 2NO₂(g).
- 2lCl(g) ↔ I₂(g) + Cl₂(g),
- Fe₂O₃(s) + 3H₂(g) ↔ 2Fe(s) + 3H₂O(g)
- 2CO(g) + 4H₂O(g) ↔ 3O₂(g) + 2CH₄(g)
- 6. $P_d(s) + 6NO(g) \leftrightarrow P_dO_d(s) + 3N_2(g)$

15.2 Calculating Equilibrium Constants

Be sure to write out the equilibrium (K.,) expression for each problem before trying to solve it!

1. What is the equilibrium expression for the following reaction?

$$SmO_2(s) + 2CO(g) \leftrightarrow Sm(s) + 2CO_2(g)$$

- At equilibrium the following concentrations are recorded: 20.0 M of H₂, 18.0 M of CO₂, 12.0 M H₂O, and 5.9 M of CO at 427 °C. What is the value of K₆₀₀ at this temperature for the following reaction? CO₂(g) + H₂(g) → CO(g) + H₂O(g)
- At equilibrium a mixture of N₂, H₂, and NH₃ gas at 500°C is determined to consist of 0.602 M N₂, 0.420 M H₂, and 0.113 M of NH₃. What is the equilibrium constant at this temperature? Does this equilibrium favor the reactants or the products? N₂(g) + 3H₂(g) ← 2NH₃(g)
- Given the equilibrium concentrations of 0.15 M NO, 0.12 M H₂ and 1.6 M N₂, Calculate the equilibrium constant for the following reaction. Does this reaction favor the reactants or the products?
 2NO(g) = 2H₂(g) ↔ N₂(g) = 2H₂O(l)
- What is the equilibrium constant for the following reaction if the equilibrium concentrations for the gases are 0.3 M for C₂H₆, 0.6 M for O₂, 1.3 M for CO₂ and 1.8 M for H₂0? Does the reaction favor the reactions or products?
 2C₂H₆(g) + 7O₂(g) 4CO₂(g) + 6H₂O(g)
- Nitrogen dioxide reacts to form dinitrogen tetraoxide. What is the equilibrium concentration of N₂O_n if the K_{eq} = 4.70 and the equilibrium concentration of NO₂ is 0.0310 M?
 2NO₂(g) → N₂O₄(g)

Calculating Equilibrium Constants Chem Worksheet 18 3

Michael Brown

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Russian Journal of Inorganic Chemistry ,1997 Geochemistry International ,1998 Vols for 1964 v 2 no 1 1965 include selected articles translated from geochemical papers from other languages but primarily from Russian German French and Chemical Equilibrium William Guenther, 1975-12 The present work is designed to provide a practical Japanese introduction to aqueous equilibrium phenomena for both students and research workers in chemistry biochemistry geochemistry and interdisciplin ary environmental fields The pedagogical strategy I have adopted makes heavy use of detailed examples of problem solving from real cases arising both in laboratory research and in the study of systems occurring in nature The procedure starts with mathematically complete equations that will provide valid solutions of equilibrium problems instead of the traditional approach through approximate concentrations and idealized infinite dilution assumptions There is repeated emphasis on the use of corrected conditional equilibrium constants and on the checking of numerical results by substitution in complete equations and or against graphs of species distributions Graphical methods of calculation and display are used extensively because of their value in clarifying equilibria and in leading one quickly to valid numerical approximations. The coverage of solution equilibrium phenomena is not however exhaustively comprehensive Rather I have chosen to offer funda mental and rigorous examinations of homogeneous step equilibria and their interactions with solubility and redox equilibria Many examples are worked out in detail to demonstrate the use of equilibrium calculations and diagrams in various fields of investigation An Analytical Investigation of Three General Methods of Calculating Chemical-equilibrium Compositions Frank J. Zeleznik, Sanford Gordon, 1960 The Brinkley Huff and White methods for chemical equilibrium calculations were modified and extended in order to permit an analytical comparison The extended forms of these methods permit condensed species as reaction products include temperature as a variable in the iteration and permit arbitrary estimates for the variables It is analytically shown that the three extended methods can be placed in a form that is independent of components In this form the Brinkley iteration is identical computationally to the White method while the modified Huff method differs only slightly from these two The convergence rates of the modified Brinkley and White methods are identical and further all three methods are guaranteed to converge and will ultimately converge quadratically It is concluded that no one of the three methods offers any significant computational advantages over the other two

Chemical Equilibrium In a Nutshell Brian Coleman,2018-06-25 Understanding the math and minutiae of chemical equilibrium can be a tall task for anyone so why not enlist the help of a scientific squirrel to guide you on your journey Join Dr Wash as we dabble in equilibrium constants and other tools needed to predict chemical processes This book focuses on introductory concepts at the high school and early university level focusing on identifying equilibrium calculating K and Q discussing Le Chatelier's principle and tying equilibrium with the field of Thermodynamics Full of step by step instructions and practice questions this book aims to simplify one of the more complex topics found within the field of chemistry

Unified Equilibrium Calculations William Benton Guenther,1991-08-19 A new approach to the academic treatment of solution equilibria is presented The author unifies homonuclear equilibrium calculations in one concept The alpha species fraction and bound proton and bound ligand ratio as a function of a single master variable the unbound H or L yield complete balances A single logic is maintained for all cases by equating the chemical binding expressed as an equilibrium condition and as a material balance condition *Equilibrium* Thomas R. Blackburn,1969 <u>Equilibrium Calculations</u> John T.

Donoghue,1971 The Determination of Stability Constants Francis J. C. Rossotti,Hazel Rossotti,1961 <u>Chemical Calculations</u> Paul Yates,1997-06-05 Chemical Calculations provides an introduction to the mathematics required for physical chemistry courses This book is unique in that it provides a gentle introduction with a chemistry centered rather than math centered approach Written by a chemist for undergraduate students it imparts an understanding of the subject from a chemist s viewpoint using examples from real chemistry It includes illustrations that show exactly how to use calculators to work problems and examples of important chemical problems with fully worked solutions This book is an ideal companion throughout a chemistry course that can be consulted when required and used to keep one step ahead of the lecture

Study Guide for Chemical Principles Thomas Elliott Taylor,1979 Equilibrium Constants of Liquid-liquid

Distribution Reactions International union of pure and applied chemistry. Commission on equilibrium data, Chemical

Equilibria in Solution M. J. Blandamer,1992 Ion Exchange Equilibrium Constants Y. Marcus,Darryl G.

Howery,International Union of Pure and Applied Chemistry. Commission on Equilibrium Data,1975 Ion Exchange Equilibrium

Constants focuses on the test compilation of equilibrium constants for ion exchange reactions The book first underscores the scope of the compilation equilibrium constants symbols used and arrangement of the table The manuscript then presents the table of equilibrium constants including polystyrene sulfonate cation exchanger polyacrylate cation exchanger polymethacrylate cation exchanger polysterene phosphate cation exchanger and zirconium phosphate cation exchanger The text highlights zirconium oxide anion exchanger zeolite type 13Y cation exchanger and

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Constant Using Spectrophotometry Norman J. Hudak, 1988-01-01

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