

Name _____

Section _____

Experiment 3**Data and Calculations: Determination of the Equilibrium Constant for a Chemical Reaction**

Trials	Volume of 0.1M Fe^{3+} (mL)	Volume of 0.1M SCN^{-} (mL)	Volume of 0.1M H^{+} (mL)	Initial $[\text{Fe}^{3+}]$ (M)	Initial $[\text{SCN}^{-}]$ (M)		Equilibrium $[\text{FeSCN}^{2+}]$ (M)
	Initial	Initial	Initial	Initial	Initial	Initial	
1	5.00	5.00	2.00	0.010	0.010	0.002	0.008
2	5.00	5.00	3.00	0.010	0.010	0.003	0.007
3	5.00	5.00	2.00	0.010	0.010	0.002	0.008
4	5.00	5.00	1.00	0.010	0.010	0.001	0.009
5	5.00	5.00	0.00	0.010	0.010	0.000	0.010

If Method B has been used, Section 3.1.1.1 of the lab manual should be followed to determine the equilibrium constant.

Equilibrium constant, K_c

Equilibrium constant, K_c

Processing the Data

A. Calculation of K_c assuming the reaction:

$\text{Fe}^{3+} + \text{SCN}^{-} \rightleftharpoons \text{FeSCN}^{2+}$

This calculation is more easily done by knowing how to change K_c to the reaction $\text{FeSCN}^{2+} \rightleftharpoons \text{Fe}^{3+} + \text{SCN}^{-}$ if you are using Method A or the lab manual has been used before the reaction was written.

Step 1: Find the initial number of moles of Fe^{3+} and SCN^{-} in the mixture (in mL) (from 3.1.1.1 of the lab manual) if you are using Method A or the lab manual has been used before.

Step 2: From the experimentally determined value of $[\text{FeSCN}^{2+}]$ at equilibrium (in mL) of the mixture, determine the value of $[\text{Fe}^{3+}]$ and $[\text{SCN}^{-}]$ at equilibrium (in mL) of the mixture. Use the value of $[\text{FeSCN}^{2+}]$ at equilibrium (in mL) of the mixture to find the number of moles of Fe^{3+} and SCN^{-} at equilibrium (in mL) of the mixture.

Equilibrium constant, K_c

Equilibrium constant, K_c

Calculations With A Chemical Reaction Lab Answers

Emilie Sanchez



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Physical Science Concepts in Action helps students make the important connection between the science they read and what they experience every day Relevant content lively explorations and a wealth of hands on activities take students understanding of science beyond the page and into the world around them Now includes even more technology tools and activities to support differentiated instruction **Innovative Education Technologies for 21st Century Teaching and Learning** Muhammad Mujtaba Asad,Fahad Sherwani,Razali Bin Hassan,Prathamesh Churi,2021-11-04 This book highlights all aspects of innovative 21st century education technologies and skills which can enhance the teaching and learning process on a broader spectrum based on best practices around the globe It offers case studies on real problems involving higher education it includes policies that need to be adaptable to the new environments such as the role of accreditation online learning MOOCs and mobile based learning The book covers all aspects of the digital competencies of teachers to fulfill the required needs of 21st century classrooms and uses a new pedagogical approach suitable for educational policies Innovative Education Technologies for 21st Teaching and Learning is the first book that addresses the teaching and learning challenges and how those challenges can be mitigated by technology which educational institutions are facing due to the COVID 19 pandemic This book is suitable for teachers students instructional and course designers policymakers and anyone interested in 21st century education **U.S. Government Research Reports** ,1963 Journal of Chemical Education ,1928
Chemical Reaction Theory Royal Society of Chemistry,1998 Papers presenting the theory of chemical reactions and the comparison with experiment **U.S. Government Research & Development Reports** ,1970 *General Catalogue*
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the Department which includes corporate securities registrations a list of tenders received and general financial data
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