

Name: \_\_\_\_\_

## Balancing Chemical Equations

Directions: Balance each of the following chemical equations.



- 1)    \_\_\_  $\text{H}_3\text{PO}_4$  + \_\_\_  $\text{KOH}$   $\rightarrow$  \_\_\_  $\text{K}_3\text{PO}_4$  + \_\_\_  $\text{H}_2\text{O}$
- 2)    \_\_\_  $\text{K}$  + \_\_\_  $\text{B}_2\text{O}_3$   $\rightarrow$  \_\_\_  $\text{K}_2\text{O}$  + \_\_\_  $\text{B}$
- 3)    \_\_\_  $\text{HCl}$  + \_\_\_  $\text{NaOH}$   $\rightarrow$  \_\_\_  $\text{NaCl}$  + \_\_\_  $\text{H}_2\text{O}$
- 4)    \_\_\_  $\text{Na}$  + \_\_\_  $\text{NaNO}_3$   $\rightarrow$  \_\_\_  $\text{Na}_2\text{O}$  + \_\_\_  $\text{N}_2$
- 5)    \_\_\_  $\text{C}$  + \_\_\_  $\text{S}_8$   $\rightarrow$  \_\_\_  $\text{CS}_2$
- 6)    \_\_\_  $\text{Na}$  + \_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_  $\text{Na}_2\text{O}$
- 7)    \_\_\_  $\text{N}_2$  + \_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_  $\text{N}_2\text{O}_5$
- 8)    \_\_\_  $\text{H}_3\text{PO}_4$  + \_\_\_  $\text{Mg}(\text{OH})_2$   $\rightarrow$  \_\_\_  $\text{Mg}_3(\text{PO}_4)_2$  + \_\_\_  $\text{H}_2\text{O}$
- 9)    \_\_\_  $\text{NaOH}$  + \_\_\_  $\text{H}_2\text{CO}_3$   $\rightarrow$  \_\_\_  $\text{Na}_2\text{CO}_3$  + \_\_\_  $\text{H}_2\text{O}$
- 10)   \_\_\_  $\text{KOH}$  + \_\_\_  $\text{HBr}$   $\rightarrow$  \_\_\_  $\text{KBr}$  + \_\_\_  $\text{H}_2\text{O}$
- 11)   \_\_\_  $\text{Na}$  + \_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_  $\text{Na}_2\text{O}$
- 12)   \_\_\_  $\text{Al}(\text{OH})_3$  + \_\_\_  $\text{H}_2\text{CO}_3$   $\rightarrow$  \_\_\_  $\text{Al}_2(\text{CO}_3)_3$  + \_\_\_  $\text{H}_2\text{O}$
- 13)   \_\_\_  $\text{Al}$  + \_\_\_  $\text{S}_8$   $\rightarrow$  \_\_\_  $\text{Al}_2\text{S}_3$
- 14)   \_\_\_  $\text{Cs}$  + \_\_\_  $\text{N}_2$   $\rightarrow$  \_\_\_  $\text{Cs}_3\text{N}$
- 15)   \_\_\_  $\text{Mg}$  + \_\_\_  $\text{Cl}_2$   $\rightarrow$  \_\_\_  $\text{MgCl}_2$
- 16)   \_\_\_  $\text{Rb}$  + \_\_\_  $\text{RbNO}_3$   $\rightarrow$  \_\_\_  $\text{Rb}_2\text{O}$  + \_\_\_  $\text{N}_2$
- 17)   \_\_\_  $\text{C}_6\text{H}_6$  + \_\_\_  $\text{O}_2$   $\rightarrow$  \_\_\_  $\text{CO}_2$  + \_\_\_  $\text{H}_2\text{O}$
- 18)   \_\_\_  $\text{N}_2$  + \_\_\_  $\text{H}_2$   $\rightarrow$  \_\_\_  $\text{NH}_3$

# Balancing Chemical Equations Kit Answers

**Cathy Duffy**



## **Balancing Chemical Equations Kit Answers:**

**X-kit Fet G11 Phys Science Chemist** Clive Long,2010      **Balancing Chemical Equations Workbook** Kraaya Publishing,2021-02-21 Chemical Reactions to Balance WorkbookThis chemistry balancing equations practice workbook contains 250 non balanced chemical equations Begin with 2 terms problems Work your way up to 6 terms problems This is the perfect workbook to increase chemistry balancing skills for beginners Table of contents How To Balance A Chemical Equation Chemical Equations To Balance Correct Answers Book features Non repetitive equations Include all reactions types synthesis combustion decomposition Use it now and develop instant recall of balancing equations Enjoy the challenge

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in reporting the answer Explains how to approach the calculations when using neat liquids or solutions The culmination of this text is the use of the same thought processes to plan the amounts of reactants needed for syntheses of desired quantities of product All of the problems in the book include detailed solutions with accompanying text to explain the answers and ancillaries also include suggestions for further reading **The Science Teacher** ,2002 SCC Library has 1964 cur

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