

Balancing Chemical Equations

Once we have written a chemical equation, we must make sure to balance it! Balancing a chemical equation means to make sure that there are the same number of each type of atom on both sides of the chemical equation. So, if there are five carbons on the left side (reactants), there needs to be five carbons on the right side (products).

So, how do we balance an equation? Follow the following steps...

- Change **only** stoichiometric coefficients – those big numbers in front of the compounds (not subscripts in the formulas – that would change the identity of the compound!)
- Start with most complex
- If found in more than one place on each side, do it last
- Watch out for odd numbers (make them into even ones if needed)
- Check your work!

Now, try to balance the following equations. In addition, name the bolded compounds for some nomenclature practice!

- $\underline{\quad} \text{NH}_4\text{NO}_3(\text{s}) \rightarrow \underline{\quad} \text{N}_2\text{O}(\text{g}) + \underline{2} \text{H}_2\text{O}(\text{g})$
ammonium nitrate & dinitrogen monoxide
- $\underline{\quad} \text{CH}_3\text{CO}_2\text{H}(\text{aq}) + \underline{2} \text{O}_2(\text{g}) \rightarrow \underline{2} \text{CO}_2(\text{g}) + \underline{2} \text{H}_2\text{O}(\text{l})$
oxygen & carbon dioxide
- $\underline{2} \text{Pb}(\text{NO}_3)_2(\text{s}) \rightarrow \underline{2} \text{N}_2\text{O}_4(\text{g}) + \underline{2} \text{PbO}(\text{s}) + \underline{\quad} \text{O}_2(\text{g})$
lead(II) nitrate & dinitrogen tetroxide & lead(II) oxide
- $\underline{\quad} \text{SiO}_2(\text{s}) + \underline{4} \text{HF} \rightarrow \underline{\quad} \text{SiF}_4(\text{g}) + \underline{2} \text{H}_2\text{O}(\text{l})$
silicon dioxide & silicon tetrafluoride
- $\underline{8} \text{H}_2\text{S}(\text{aq}) + \underline{8} \text{Cl}_2(\text{g}) \rightarrow \underline{\quad} \text{S}_8(\text{s}) + \underline{16} \text{HCl}(\text{aq})$
chlorine
- $\underline{\quad} \text{H}_2\text{CO}_3(\text{aq}) + \underline{\quad} \text{Ba}(\text{OH})_2 \rightarrow \underline{\quad} \text{BaCO}_3(\text{s}) + \underline{2} \text{H}_2\text{O}(\text{l})$
barium hydroxide & barium carbonate
- $\underline{\quad} \text{P}_4(\text{s}) + \underline{6} \text{Br}_2(\text{l}) \rightarrow \underline{4} \text{PBr}_3(\text{l})$
phosphorous & phosphorous tribromide
- $\underline{\quad} \text{S}_8(\text{l}) + \underline{4} \text{Cl}_2(\text{g}) \rightarrow \underline{4} \text{S}_2\text{Cl}_2(\text{g})$
disulfur dichloride
- $\underline{2} \text{B}_5\text{H}_9(\text{l}) + \underline{12} \text{O}_2(\text{g}) \rightarrow \underline{5} \text{B}_2\text{O}_3(\text{s}) + \underline{9} \text{H}_2\text{O}(\text{g})$
pentaboron nonahydride & diboron trioxide
- $\underline{3} \text{H}_2\text{SO}_4(\text{aq}) + \underline{2} \text{Al}(\text{OH})_3(\text{s}) \rightarrow \underline{\quad} \text{Al}_2(\text{SO}_4)_3(\text{aq}) + \underline{6} \text{H}_2\text{O}(\text{l})$
aluminum hydroxide & aluminum sulfate
- $\underline{2} \text{Al}(\text{s}) + \underline{3} \text{Br}_2(\text{l}) \rightarrow \underline{2} \text{AlBr}_3(\text{s})$
aluminum bromide
- $\underline{\quad} (\text{NH}_4)_2\text{Cr}_2\text{O}_7(\text{s}) \rightarrow \underline{\quad} \text{N}_2(\text{g}) + \underline{\quad} \text{Cr}_2\text{O}_3(\text{s}) + \underline{4} \text{H}_2\text{O}(\text{l})$
ammonium dichromate & chromium(III) oxide
- $\underline{\quad} \text{S}_8(\text{s}) + \underline{24} \text{F}_2(\text{g}) \rightarrow \underline{8} \text{SF}_6(\text{g})$
sulfur hexafluoride
- $\underline{\quad} \text{Zn}(\text{s}) + \underline{\quad} \text{Pb}(\text{NO}_3)_2(\text{aq}) \rightarrow \underline{\quad} \text{Zn}(\text{NO}_3)_2(\text{aq}) + \underline{\quad} \text{Pb}(\text{s})$
zinc nitrate

Balancing Chemical Equations Multiple Choice Test

Richard H. Langley, John Moore



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Table of Contents Balancing Chemical Equations Multiple Choice Test

1. Understanding the eBook Balancing Chemical Equations Multiple Choice Test
 - The Rise of Digital Reading Balancing Chemical Equations Multiple Choice Test
 - Advantages of eBooks Over Traditional Books
2. Identifying Balancing Chemical Equations Multiple Choice Test
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Balancing Chemical Equations Multiple Choice Test
 - User-Friendly Interface
4. Exploring eBook Recommendations from Balancing Chemical Equations Multiple Choice Test
 - Personalized Recommendations
 - Balancing Chemical Equations Multiple Choice Test User Reviews and Ratings
 - Balancing Chemical Equations Multiple Choice Test and Bestseller Lists
5. Accessing Balancing Chemical Equations Multiple Choice Test Free and Paid eBooks
 - Balancing Chemical Equations Multiple Choice Test Public Domain eBooks
 - Balancing Chemical Equations Multiple Choice Test eBook Subscription Services
 - Balancing Chemical Equations Multiple Choice Test Budget-Friendly Options
6. Navigating Balancing Chemical Equations Multiple Choice Test eBook Formats

- ePub, PDF, MOBI, and More
 - Balancing Chemical Equations Multiple Choice Test Compatibility with Devices
 - Balancing Chemical Equations Multiple Choice Test Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Balancing Chemical Equations Multiple Choice Test
 - Highlighting and Note-Taking Balancing Chemical Equations Multiple Choice Test
 - Interactive Elements Balancing Chemical Equations Multiple Choice Test
 8. Staying Engaged with Balancing Chemical Equations Multiple Choice Test
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Balancing Chemical Equations Multiple Choice Test
 9. Balancing eBooks and Physical Books Balancing Chemical Equations Multiple Choice Test
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Balancing Chemical Equations Multiple Choice Test
 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
 11. Cultivating a Reading Routine Balancing Chemical Equations Multiple Choice Test
 - Setting Reading Goals Balancing Chemical Equations Multiple Choice Test
 - Carving Out Dedicated Reading Time
 12. Sourcing Reliable Information of Balancing Chemical Equations Multiple Choice Test
 - Fact-Checking eBook Content of Balancing Chemical Equations Multiple Choice Test
 - Distinguishing Credible Sources
 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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