

Name: _____
Hour: _____

Calculating Thermal Energy Specific Heat Worksheet #2

1. What is the specific heat of a substance that absorbs 2500 joules of heat when a sample of 1200 g of the substance increases in temperature from 10.0°C to 70.0°C ?
2. How many grams of water would require 2250 joules of heat to raise its temperature from 34.0°C to 100.0°C ?
3. If 200 g of water is to be heated from 24.0°C to 100.0°C to make a cup of tea, how much heat must be added?
4. A block of aluminum weighing 140 g is cooled from 98.4°C to 62.2°C with the release of 1080 joules of heat. From this data, calculate the specific heat of aluminum.
5. A cube of gold weighing 192.4g is heated from 30.0°C to some higher temperature, with the absorption of 226 joules of heat. The specific heat of gold is $0.030 \text{ J/g}\cdot^{\circ}\text{C}$. What was the final temperature of the gold?
6. A total of 54 joules of heat are absorbed as 58.3 g of lead is heated from 12°C to 42°C . From this data, what is the specific heat of lead?
7. The specific heat of wood is $2.03 \text{ J/g}\cdot^{\circ}\text{C}$. How much heat is needed to convert 550 g of wood at -15.0°C to 10.0°C into energy?
8. Granite has a specific heat of $800 \text{ J/g}\cdot^{\circ}\text{C}$. What mass of granite is needed to store 150,200 J of heat if the temperature of the granite is to be increased by 15.5°C ?

Calculating Specific Heat Answer Key

Marian DeWane, Heather Hattori

A red circular graphic with a gradient, appearing as a partial circle or a stylized arrow pointing to the right, located to the right of the authors' names.

Calculating Specific Heat Answer Key:

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