

# Stoichiometry

## Section 11.1 Defining Stoichiometry

pages 368–372

### Practice Problems

pages 371–372

1. Interpret the following balanced chemical equations in terms of particles, moles, and mass. Show that the law of conservation of mass is observed.



1 molecule  $\text{N}_2$  + 3 molecules

$\text{H}_2 \rightarrow$  2 molecules  $\text{NH}_3$

1 mole  $\text{N}_2$  + 3 moles  $\text{H}_2 \rightarrow$  2 moles  $\text{NH}_3$

Mass:

$\text{N}_2$ : (2 mol)(14.007 g/mol) = 28.014 g

$3\text{H}_2$ : (6 mol)(1.008 g/mol) = 6.048 g

$2\text{NH}_3$ : (2 mol)(14.007 g/mol) +  
(6 mol)(1.008 g/mol) = 34.062 g

28.014 g  $\text{N}_2$  + 6.048 g  $\text{H}_2 \rightarrow$  34.062 g  $\text{NH}_3$

34.062 g reactants = 34.062 g products



1 molecule  $\text{HCl}$  + 1 formula unit  $\text{KOH} \rightarrow$

1 formula unit  $\text{KCl}$  + 1 molecule  $\text{H}_2\text{O}$

1 mole  $\text{HCl}$  + 1 mole  $\text{KOH} \rightarrow$

1 mole  $\text{KCl}$  + 1 mole  $\text{H}_2\text{O}$

Mass:

$\text{HCl}$ : (1 mol)(1.008 g/mol) +  
(1 mol)(35.453 g/mol) = 36.461 g

$\text{KOH}$ : (1 mol)(39.098 g/mol) +  
(1 mol)(15.999 g/mol) + (1 mol)(1.008 g/mol)  
= 56.105 g

$\text{KCl}$ : (1 mol)(39.098 g/mol) +  
(1 mol)(35.453 g/mol) = 74.551 g

$\text{H}_2\text{O}$ : (2 mol)(1.008 g/mol) +  
(1 mol)(15.999 g/mol) = 18.015 g

36.461 g  $\text{HCl}$  + 56.105 g  $\text{KOH} \rightarrow$   
74.551 g  $\text{KCl}$  + 18.015 g  $\text{H}_2\text{O}$

92.566 g reactants = 92.566 g products



2 atoms  $\text{Mg}$  + 1 molecule

$\text{O}_2 \rightarrow$  2 formula units  $\text{MgO}$

2 moles  $\text{Mg}$  + 1 mole  $\text{O}_2 \rightarrow$  2 moles  $\text{MgO}$

Mass:

$2\text{Mg}$ : (2 mol)(24.305 g/mol) = 48.610 g

$\text{O}_2$ : (2 mol)(15.999 g/mol) = 31.998 g

$2\text{MgO}$ : (2 mol)(24.305 g/mol) +  
(2 mol)(15.999 g/mol) = 80.608 g

48.610 g  $\text{Mg}$  + 31.998 g  $\text{O}_2 \rightarrow$  80.608 g  $\text{MgO}$

80.608 g reactants = 80.608 g products

2. **Challenge** For each of the following, balance the chemical equation; interpret the equation in terms of particles, moles, and mass; and show that the law of conservation of mass is observed.



2 atoms  $\text{Na}$  + 2 molecules  $\text{H}_2\text{O} \rightarrow$

2 formula units  $\text{NaOH}$  + 1 molecule  $\text{H}_2$

2 mol  $\text{Na}$  + 2 mol  $\text{H}_2\text{O} \rightarrow$

2 mol  $\text{NaOH}$  + 1 mol  $\text{H}_2$

Mass:

$2\text{Na}$ : (2 mol)(22.990 g/mol) = 45.980 g

$2\text{H}_2\text{O}$ : (4 mol)(1.008 g/mol) +  
(2 mol)(15.999 g/mol) = 36.030 g

$2\text{NaOH}$ : (2 mol)(22.990 g/mol) +  
(2 mol)(15.999 g/mol) + (2 mol)(1.008 g/mol)  
= 79.994 g

$\text{H}_2$ : (2 mol)(1.008 g/mol) = 2.016 g

45.980 g  $\text{Na}$  + 36.030 g  $\text{H}_2\text{O} \rightarrow$

79.994 g  $\text{NaOH}$  + 2.016 g  $\text{H}_2$

82.01 g reactants = 82.01 g products

# Chapter 11 Stoichiometry Study Guide Answers

**Bernice G. Segal, Peter S. Shenkin**



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## **Table of Contents Chapter 11 Stoichiometry Study Guide Answers**

1. Understanding the eBook Chapter 11 Stoichiometry Study Guide Answers
  - The Rise of Digital Reading Chapter 11 Stoichiometry Study Guide Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Chapter 11 Stoichiometry Study Guide Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 11 Stoichiometry Study Guide Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Chapter 11 Stoichiometry Study Guide Answers
  - Personalized Recommendations
  - Chapter 11 Stoichiometry Study Guide Answers User Reviews and Ratings
  - Chapter 11 Stoichiometry Study Guide Answers and Bestseller Lists
5. Accessing Chapter 11 Stoichiometry Study Guide Answers Free and Paid eBooks

- Chapter 11 Stoichiometry Study Guide Answers Public Domain eBooks
- Chapter 11 Stoichiometry Study Guide Answers eBook Subscription Services
- Chapter 11 Stoichiometry Study Guide Answers Budget-Friendly Options
- 6. Navigating Chapter 11 Stoichiometry Study Guide Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Chapter 11 Stoichiometry Study Guide Answers Compatibility with Devices
  - Chapter 11 Stoichiometry Study Guide Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 11 Stoichiometry Study Guide Answers
  - Highlighting and Note-Taking Chapter 11 Stoichiometry Study Guide Answers
  - Interactive Elements Chapter 11 Stoichiometry Study Guide Answers
- 8. Staying Engaged with Chapter 11 Stoichiometry Study Guide Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 11 Stoichiometry Study Guide Answers
- 9. Balancing eBooks and Physical Books Chapter 11 Stoichiometry Study Guide Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 11 Stoichiometry Study Guide Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 11 Stoichiometry Study Guide Answers
  - Setting Reading Goals Chapter 11 Stoichiometry Study Guide Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 11 Stoichiometry Study Guide Answers
  - Fact-Checking eBook Content of Chapter 11 Stoichiometry Study Guide Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development

- Exploring Educational eBooks

#### 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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