#### CHAPTER 19 REVIEW

## Oxidation-Reduction Reactions

## SECTION 1

SHORT ANSWER Answer the following questions in the space provided.

- All the following equations involve redex reactions except
  - (a) CuO + H<sub>2</sub>O → Ca(OH)<sub>2</sub>.
  - (b) 2SO<sub>2</sub> + O<sub>2</sub> → 2SO<sub>3</sub>.
  - (c)  $2HgO \rightarrow 2Hg + O_2$ .
  - (d) SnCl<sub>4</sub> + 2FeCl<sub>2</sub> → 2FeCl<sub>3</sub> + SnCl<sub>2</sub>.
- Assign the correct exidation number to the individual atom or ion below.
  - \_\_\_\_\_ a. Ma in MaO<sub>2</sub>
  - \_\_\_\_\_ b. S in S<sub>R</sub>
  - e. Cl in CaCl<sub>2</sub>
  - \_\_\_\_\_\_ d. 1 is 10<sub>3</sub>
  - \_\_\_\_\_\_ e. C in HCO<sub>3</sub>
  - f. Fe in Fe<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub>
  - \_\_\_\_\_ g. S in Fe<sub>2</sub>(SO<sub>4</sub>)<sub>3</sub>
- 3. In each of the following half-reactions, determine the value of x.
  - \_\_\_\_\_ a.  $S^{6+} + x e^- \rightarrow S^{2-}$
  - \_\_\_\_\_ b.  $2Br'' \rightarrow Br_2 + 2e^-$
  - \_\_\_\_\_ c.  $Sn^{4+} + 2e^- \rightarrow Sn^a$
  - d. Which of the above half-reactions represent reduction processes?
- 4. Give examples, other than those listed in Table 1 on page 631 of the text, for the following:
  - \_\_\_\_\_\_ n. a compound containing H in a −1 oxidation state
  - \_\_\_\_\_\_ b. a peroxide
  - e. a polyatomic son in which the oxidation number for S is +4
  - d. a substance in which the excidation number for F is not −1

**Becker, Charles F. Batten** 

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#### **Table of Contents Chapter 19 Practice Problems Oxidation Reduction Reactions Answers**

- 1. Understanding the eBook Chapter 19 Practice Problems Oxidation Reduction Reactions Answers
  - The Rise of Digital Reading Chapter 19 Practice Problems Oxidation Reduction Reactions Answers
  - Advantages of eBooks Over Traditional Books
- 2. Identifying Chapter 19 Practice Problems Oxidation Reduction Reactions Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
- 3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Chapter 19 Practice Problems Oxidation Reduction Reactions Answers
  - User-Friendly Interface
- 4. Exploring eBook Recommendations from Chapter 19 Practice Problems Oxidation Reduction Reactions Answers
  - Personalized Recommendations
  - Chapter 19 Practice Problems Oxidation Reduction Reactions Answers User Reviews and Ratings
  - Chapter 19 Practice Problems Oxidation Reduction Reactions Answers and Bestseller Lists
- 5. Accessing Chapter 19 Practice Problems Oxidation Reduction Reactions Answers Free and Paid eBooks
  - o Chapter 19 Practice Problems Oxidation Reduction Reactions Answers Public Domain eBooks
  - o Chapter 19 Practice Problems Oxidation Reduction Reactions Answers eBook Subscription Services
  - o Chapter 19 Practice Problems Oxidation Reduction Reactions Answers Budget-Friendly Options
- 6. Navigating Chapter 19 Practice Problems Oxidation Reduction Reactions Answers eBook Formats
  - o ePub, PDF, MOBI, and More
  - $\circ$  Chapter 19 Practice Problems Oxidation Reduction Reactions Answers Compatibility with Devices
  - Chapter 19 Practice Problems Oxidation Reduction Reactions Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Chapter 19 Practice Problems Oxidation Reduction Reactions Answers
  - Highlighting and Note-Taking Chapter 19 Practice Problems Oxidation Reduction Reactions Answers
  - Interactive Elements Chapter 19 Practice Problems Oxidation Reduction Reactions Answers

- 8. Staying Engaged with Chapter 19 Practice Problems Oxidation Reduction Reactions Answers
  - Joining Online Reading Communities
  - o Participating in Virtual Book Clubs
  - Following Authors and Publishers Chapter 19 Practice Problems Oxidation Reduction Reactions Answers
- 9. Balancing eBooks and Physical Books Chapter 19 Practice Problems Oxidation Reduction Reactions Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Chapter 19 Practice Problems Oxidation Reduction Reactions Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Chapter 19 Practice Problems Oxidation Reduction Reactions Answers
  - Setting Reading Goals Chapter 19 Practice Problems Oxidation Reduction Reactions Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Chapter 19 Practice Problems Oxidation Reduction Reactions Answers
  - Fact-Checking eBook Content of Chapter 19 Practice Problems Oxidation Reduction Reactions Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends
  - $\circ \ \ Integration \ of \ Multimedia \ Elements$
  - Interactive and Gamified eBooks

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