

Science

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(Chapter – 3) (Atoms and Molecules) (Exercises)
(Class – IX)

Question 1:

A 0.24 g sample of compound of oxygen and boron was found by analysis to contain 0.096 g of boron and 0.144 g of oxygen. Calculate the percentage composition of the compound by weight.

Answer 1:

Mass of boron = 0.096g

(Given) Mass of oxygen = 0.144g (Given) Mass of sample = 0.24g (Given)

Thus, percentage of boron by weight in the compound = $\frac{0.096 \times 100}{0.24} \% = 40\%$

Thus, percentage of oxygen by weight in the compound = $\frac{0.144 \times 100}{0.24} \% = 60\%$

Question 2:

When 3.0 g of carbon is burnt in 8.00 g oxygen, 11.00 g of carbon dioxide is produced. What mass of carbon dioxide will be formed when 3.00 g of carbon is burnt in 50.00 g of oxygen? Which law of chemical combination will govern your answer?

Answer 2:

Carbon + Oxygen → Carbon dioxide 3g of carbon reacts with 8 g of oxygen to produce 11g of carbon dioxide. If 3g of carbon is burnt in 50g of oxygen, then 3g of carbon will react with 8 g of oxygen. The remaining 42 g of oxygen will be left un-reactive. In this case also, only 11g of carbon dioxide will be formed. The above answer is governed by the law of constant proportions.

Question 3:

What are polyatomic ions? Give examples.

Answer 3:

A polyatomic ion is a group of atoms carrying a charge (positive or negative). For example, ammonium ion (NH_4^+), hydroxide ion (OH^-), carbonate ion (CO_3^{2-}), sulphate ion (SO_4^{2-}).

Question 4:

Write the chemical formulae of the following.

- (a) Magnesium chloride
- (b) Calcium oxide
- (c) Copper nitrate
- (d) Aluminium chloride
- (e) Calcium carbonate.

Answer 4:

- (a) Magnesium chloride → $MgCl_2$
- (b) Calcium oxide → CaO
- (c) Copper nitrate → $Cu(NO_3)_2$
- (d) Aluminium chloride → $AlCl_3$
- (e) Calcium carbonate → $CaCO_3$

Question 5:

Write the chemical formulae of the following.

- (a) Quick lime
- (b) Hydrogen bromide
- (c) Baking powder
- (d) Potassium sulphate.

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


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















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