

## Types of Chemical Reactions Lab Answer Key

### Chemical Reaction #1

|                                              |                                                                                                                      |
|----------------------------------------------|----------------------------------------------------------------------------------------------------------------------|
| <i>Reactants<br/>(write in words)</i>        | hydrochloric acid and magnesium                                                                                      |
| <i>Observations</i>                          | accept most responses                                                                                                |
| <i>Type of Reaction</i>                      | single replacement                                                                                                   |
| <i>Predict Products<br/>(write in words)</i> | magnesium chloride and hydrogen gas                                                                                  |
| <i>Equation<br/>(write in symbols)</i>       | $2\text{HCl}_{\text{(aq)}} + \text{Mg}_{\text{(s)}} \rightarrow \text{MgCl}_{2\text{(aq)}} + \text{H}_{2\text{(g)}}$ |

### Chemical Reaction #2

|                                              |                                                                                         |
|----------------------------------------------|-----------------------------------------------------------------------------------------|
| <i>Reactants<br/>(write in words)</i>        | copper metal and oxygen gas                                                             |
| <i>Observations</i>                          | accept most responses                                                                   |
| <i>Type of Reaction</i>                      | synthesis                                                                               |
| <i>Predict Products<br/>(write in words)</i> | copper oxide                                                                            |
| <i>Equation<br/>(write in symbols)</i>       | $2\text{Cu}_{\text{(s)}} + \text{O}_{2\text{(g)}} \rightarrow 2\text{CuO}_{\text{(s)}}$ |

### Chemical Reaction #3

|                                              |                                                                                                                            |
|----------------------------------------------|----------------------------------------------------------------------------------------------------------------------------|
| <i>Reactants<br/>(write in words)</i>        | methane gas and oxygen gas                                                                                                 |
| <i>Observations</i>                          | accept most responses                                                                                                      |
| <i>Type of Reaction</i>                      | combustion                                                                                                                 |
| <i>Predict Products<br/>(write in words)</i> | carbon dioxide and water                                                                                                   |
| <i>Equation<br/>(write in symbols)</i>       | $\text{CH}_{4\text{(g)}} + 2\text{O}_{2\text{(g)}} \rightarrow \text{CO}_{2\text{(g)}} + 2\text{H}_2\text{O}_{\text{(g)}}$ |

### Chemical Reaction #4

|                                              |                                                                                                           |
|----------------------------------------------|-----------------------------------------------------------------------------------------------------------|
| <i>Reactants<br/>(write in words)</i>        | hydrogen peroxide                                                                                         |
| <i>Observations</i>                          | accept most responses                                                                                     |
| <i>Type of Reaction</i>                      | decomposition                                                                                             |
| <i>Predict Products<br/>(write in words)</i> | water and oxygen gas                                                                                      |
| <i>Equation<br/>(write in symbols)</i>       | $2\text{H}_2\text{O}_{2\text{(l)}} \rightarrow 2\text{H}_2\text{O}_{\text{(l)}} + \text{O}_{2\text{(g)}}$ |

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**Cambridge**



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