

DETERMINING MOLECULAR FORMULAS (TRUE FORMULAS)

Name: _____

Derive the molecular formula:

1. The empirical formula of a compound is PCl_2 . The molecular mass is 263 g/mol. What is the molecular formula?

Correct answer: PCl_2

Score: _____

2. The empirical formula of a compound is CH_2 . The molecular mass is 78 g/mol. What is the molecular formula?

Correct answer: C_6H_6

Score: _____

3. A compound is found to contain 60.0% carbon, 4.48% hydrogen and 35.5% oxygen. The molecular mass is 168 g/mol. What is the molecular formula?

Correct answer: $\text{C}_8\text{H}_8\text{O}_2$

Score: _____

4. A compound is found to contain 60.0% carbon, 4.48% hydrogen and 35.5% oxygen. The molecular mass is 168 g/mol. What is the molecular formula?

Correct answer: $\text{C}_8\text{H}_8\text{O}_2$

Score: _____

5. A compound is found to contain 60.0% carbon, 4.48% hydrogen and 35.5% oxygen. The molecular mass is 168 g/mol. What is the molecular formula?

Correct answer: $\text{C}_8\text{H}_8\text{O}_2$

Score: _____

Determining Molecular Formula True Formulas Answer Key

B Lingard



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