

## Determining Empirical Formulas

1) 75% carbon, 25% hydrogen	Answer:	$\text{CH}_4$
Work: $\text{C } 75\% \rightarrow 75\text{g C} \left  \frac{1 \text{ mol}}{12.01 \text{ g C}} = 6.24 \text{ mol} / 6.24 = 1 \right.$ $\text{H } 25\% \rightarrow 25\text{g H} \left  \frac{1 \text{ mol H}}{1.01 \text{ g H}} = 24.75 \text{ mol} / 6.24 = 4 \right.$ <div style="float: right; border: 1px solid black; padding: 5px;"><math>\text{CH}_4</math></div>		
2) 52.7% potassium, 47.3% chlorine	Answer:	$\text{KCl}$
Work: $\text{K } 52.7\% \rightarrow 52.7\text{g K} \left  \frac{1 \text{ mol K}}{39.10 \text{ g K}} = 1.34 \text{ mol} / 1.33 = 1 \right.$ $\text{Cl } 47.3\% \rightarrow 47.3\text{g Cl} \left  \frac{1 \text{ mol Cl}}{35.45 \text{ g Cl}} = 1.33 \text{ mol} / 1.33 = 1 \right.$ <div style="float: right; border: 1px solid black; padding: 5px;"><math>\text{KCl}</math></div>		
3) 22.1% aluminum, 25.4% phosphorous, 52.5 % oxygen	Answer:	$\text{AlPO}_4$
Work: $\text{Al } 22.1\% \rightarrow 22.1\text{g Al} \left  \frac{1 \text{ mol Al}}{26.98 \text{ g Al}} = 0.819 \text{ mol} / 0.819 = 1 \right.$ $\text{P } 25.4\% \rightarrow 25.4\text{g P} \left  \frac{1 \text{ mol P}}{30.97 \text{ g P}} = 0.820 \text{ mol} / 0.819 = 1 \right.$ $\text{O } 52.5\% \rightarrow 52.5\text{g O} \left  \frac{1 \text{ mol O}}{16.00 \text{ g O}} = 3.28 \text{ mol} / 0.819 = 4 \right.$ <div style="float: right; border: 1px solid black; padding: 5px;"><math>\text{AlPO}_4</math></div>		
4) 13% magnesium, 87% bromine	Answer:	$\text{MgBr}_2$
Work: $13\text{g Mg} \left  \frac{1 \text{ mol Mg}}{24.31 \text{ g Mg}} = 0.54 \text{ mol Mg} / 0.54 = 1 \right.$ $87\text{g Br} \left  \frac{1 \text{ mol Br}}{79.90 \text{ g Br}} = 1.09 \text{ mol Br} / 0.54 = 2 \right.$ <div style="float: right; border: 1px solid black; padding: 5px;"><math>\text{MgBr}_2</math></div>		
5) 32.4% sodium, 22.5% sulfur, 45.1% oxygen	Answer:	$\text{Na}_2\text{SO}_4$
Work: $32.4\text{g Na} \left  \frac{1 \text{ mol}}{22.99 \text{ g Na}} = 1.41 \text{ mol Na} / 0.70 = 2 \right.$ $22.5\text{g S} \left  \frac{1 \text{ mol S}}{32.37 \text{ g S}} = 0.70 \text{ mol S} / 0.70 = 1 \right.$ $45.1\text{g O} \left  \frac{1 \text{ mol O}}{16 \text{ g O}} = 2.82 \text{ mol O} / 0.70 = 4 \right.$ <div style="float: right; border: 1px solid black; padding: 5px;"><math>\text{Na}_2\text{SO}_4</math></div>		

# Determining Chemical Formulas Answer Key

**Thomas Elliott Taylor**



## Determining Chemical Formulas Answer Key:

**Regents Exams and Answers: Chemistry--Physical Setting Revised Edition** Barron's Educational Series, Albert Tarendash, 2021-01-05 Barron's Regents Exams and Answers Chemistry provides essential practice for students taking the Chemistry Regents including actual recently administered exams and thorough answer explanations for all questions This book features Eight actual administered Regents Chemistry exams so students can get familiar with the test Thorough explanations for all answers Self analysis charts to help identify strengths and weaknesses Test taking techniques and strategies A detailed outline of all major topics tested on this exam A glossary of important terms to know for test day

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