

## Chemical and Physical Changes and Properties

### Purpose

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Based on experimental data, determine the occurrence of a chemical change.

### Introduction

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Text Reference: Chapter 1, *Chemistry, A Molecular Approach*, Nivaldo Tro, 4<sup>th</sup> ed.

Matter can experience two types of changes:

- A physical change is a change in physical properties (size, mass, state of matter, etc) without a change in chemical composition.
- A chemical change involves the rearrangement of atoms of one or more substances and a change in chemical properties or composition, resulting in the formation of at least one new substance. All chemical changes are accompanied by one or more physical change, along with energy exchanges.

A common example of a physical change is a change in state of the substance. The appearance of a substance changes when you melt, freeze, boil, condense or sublime it. This change is due to the addition or removal of heat. The composition of the substance remains the same, even if it looks very different. When heat is added to ice, it starts to melt and then the appearance is completely different, from a clear solid to a transparent liquid. If we continue adding heat, the liquid starts to vaporize, and the water now is steam. Even when the ice does not look very similar to steam, the composition is the same: it is water. Cutting, tearing, shattering, and grinding are other examples of physical changes. These may be irreversible, however there is no change in composition.

In a chemical change, a change in the composition occurs. For example, when the Statue of Liberty was erected in 1886, the copper metal was orange in color. However, by 1900, it had turned a turquoise color. The copper underwent a chemical change, going from metallic copper to a mixture of copper oxides, copper carbonates, copper sulfates, and copper chlorides due to exposure to various chemicals in the air and rain.

Certain observations that indicate a chemical change include:

- *Formation of precipitate.* It is the formation of an insoluble solid when two reactants are mixed. It may be a fine suspension or it may form large particles that quickly settle down. It may occur immediately, or it may take some time to settle down in order to be observed.
- *Color change.* It is the observation of a new color in the substance(s) that we are using. When a clean copper wire is heated in a burner flame, the appearance of the copper changes from coppery metallic to glowing red. Finally, its final appearance by cooling is a black solid. It is a new substance; copper oxide. However, when a clean platinum wire

# Chm2045l Lab Manual

**J. A. Beran, Mark Lassiter**



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