

## **USER'S MANUAL SMOKE ALARMS**

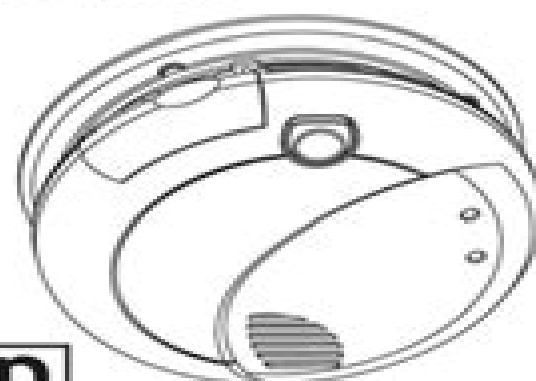
***First Alert***

### **AC Powered Photoelectric Smoke Alarms**

Input: 120V AC ~, 60Hz, 0.05A

**Model 7010B**  
with Battery Back-Up and Silence

**Model 7020B**  
with Battery Back-Up, Silence  
and Escape Light® Feature



#### **IMPORTANT! PLEASE READ CAREFULLY AND SAVE.**

This user's manual contains important information about your Smoke Alarm's operation. If you are installing this Smoke Alarm for use by others, you must leave this manual—or a copy of it—with the end user.

Printed in Mexico  
M08-0050-011 Q 11/15

**CONFORMS TO**  
**UL STD 217** Models  
7010B, 7020B

# Brk First Alert 7010b Manual

**International Municipal Signal  
Association**

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the International Municipal Signal Association text.

**Brk First Alert 7010b Manual:**

**Official Fire Alarm Manual of Installation and Maintenance Procedures** International Municipal Signal Association, 1972

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though slogans such as nursing must be a research based profession have been around for a long time recent initiatives such as nursing development units ndus quality assurance and evidence based

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web introduction this laboratory exercise relies on a titration technique to determine an unknown concentration of monoprotic acid in solution in the process of titration a basic solution is gradually added to the acidic solution until complete neutralization is obtained

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web with a strong acid and strong base titration strong acids and strong bases refer to species that completely dissociate to form ions in solution by contrast weak acids and bases ionize only partially and the ionization reaction is reversible

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web aug 30 2022 acid base titrations are usually used to find the amount of a known acidic or basic substance through acid base reactions the analyte titrand is the solution with an unknown molarity the reagent titrant is the solution with a known molarity that will react with the analyte

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web conclusion the purpose of the experiment was to identify the unknown solid acid 17 by calculating its molar mass this was done by titration of the acid with sodium hydroxide in order to find the volume of sodium hydroxide required to reach the end point and that value was used in unit analysis to find the molar mass and identify the unknown solid acid

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web by utilizing the half equivalence point on the titration curve the pka of the acid analyte was determined to be 3.655 furthermore using the stoichiometric ratio between the naoh and the acid analyte 1:1 we could experimentally determine the number of moles of the acid in the solution to be 0.00731 moles making the experimental molar mass

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web goals understand that parts of an acid base titration be able to determine the  $K_a$  or  $K_b$  from pH data associated with the

titration of a weak acid or base be able to determine the molar mass of a solid monoprotic acid from titration data be able to calculate  $K_{a1}$  and  $K_{a2}$  for a polyprotic acid by the end of this lab students should be able to

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web titration experiment in the neutralization of hydrochloric acid by sodium hydroxide the mole ratio of acid to base is 1 1  
 $\text{HCl} + \text{NaOH} \rightarrow \text{NaCl} + \text{H}_2\text{O}$   
 1 mole of HCl would be fully neutralized by one mole of NaOH

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web conclusion the identification of an unknown solid acid the purpose of this experiment was to determine the standardization of sodium hydroxide NaOH and oxalic acid  $\text{H}_2\text{C}_2\text{O}_4 \cdot 2\text{H}_2\text{O}$  and the molar mass of an unknown solid acid using the molar mass of the standardized sodium hydroxide solution

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web acid base reactions can be used to determine concentrations of an acid or a base using stoichiometric calculations and titrimetric analysis the volume of a solution required to react

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web jul 19 2023 introduction titration is an analytical quantitative technique used to determine the concentration of a solute a pH titration is used to determine the concentration of an acid or a base titrations play an important role in determining amount and purity in many manufacturing processes

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web specifically an acid base titration can be used to figure out the concentration of an acid or base whether an unknown acid or base is strong or weak  $pK_a$  of an unknown acid or  $pK_b$  of the unknown base key terms titration a process where a solution of known strength is added to a certain volume of a treated sample containing an

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web conclusion this lab evaluated an acid base reaction with the volumetric analysis of a standard solution being NaOH reaction with a 2nd substance  $\text{H}_2\text{C}_2\text{O}_4 + 2\text{NaOH} \rightarrow \text{Na}_2\text{C}_2\text{O}_4 + 2\text{H}_2\text{O}$  to differentiate an acid from a base in a solution a pH scale is used the pH scale ranges from 0 to 14 a pH of 7 is neutral in an acid base reaction

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week 1

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web this point of titration is referred to as the equivalence point considering the 1 1 stoichiometry of this acid base reaction  
 $\text{NaOH (aq)} + \text{C}_6\text{H}_4\text{COOH (aq)} \rightarrow \text{C}_6\text{H}_4\text{COONa (aq)} + \text{H}_2\text{O (l)}$   
the point of equivalence is the point of titration when the number of moles of NaOH added is equal to the number of moles of KHP in the