

Student Exploration: Covalent Bonds

Vocabulary: covalent bond, diatomic molecule, Lewis diagram, molecule, noble gases, nonmetal, octet rule, shell, valence, valence electron


Prior Knowledge Questions (Do these BEFORE using the Gizmo.)


1. There are eight markers in a full set, but Flora and Frank each only have seven markers. Flora is missing the red marker, and Frank is missing the blue marker. What can they do so that each has a full set of markers? **Frank should give the red marker to Flora, and Flora should give the blue marker to Frank.**
2. Otto and Olivia each have six markers. Otto is missing the purple and green markers, and Olivia is missing the black and brown markers. What can they do so that each has a full set? **Otto should give the black and brown markers to Olivia, and Olivia should give the purple and green markers to Otto.**

Gizmo Warm-up

Just like the students described above, **nonmetal** atoms can share electrons. As you will see in the *Covalent Bonds* Gizmo™, atoms form bonds in this way.



To begin, check that **Fluorine** is selected from the **Select a substance** menu. Click **Play** () to see the electrons orbiting the nucleus of each atom.

1. The outermost electrons in each atom are called **valence electrons**. How many valence electrons does each fluorine atom have? **7 valence electrons.**
2. Click **Pause** (). Drag an electron from the left atom to the right atom. Click **Play**.
What happens? **They equally share the one atom on the new shell that is formed.**
3. Click **Pause**, drag an electron from the right atom to the left, and then click **Play**.

What happens now? **They continue to spiral around both atoms, but now there are two electrons on the shell.**

Activity A:
Sharing electrons

Get the Gizmo ready:

- Click **Reset**.
- Select **Hydrogen**.



Covalent Bonding Exploration Guide Answers

Victor M. Corman



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Table of Contents Covalent Bonding Exploration Guide Answers

1. Understanding the eBook Covalent Bonding Exploration Guide Answers
 - The Rise of Digital Reading Covalent Bonding Exploration Guide Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Covalent Bonding Exploration Guide Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Covalent Bonding Exploration Guide Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Covalent Bonding Exploration Guide Answers
 - Personalized Recommendations
 - Covalent Bonding Exploration Guide Answers User Reviews and Ratings
 - Covalent Bonding Exploration Guide Answers and Bestseller Lists

5. Accessing Covalent Bonding Exploration Guide Answers Free and Paid eBooks
 - Covalent Bonding Exploration Guide Answers Public Domain eBooks
 - Covalent Bonding Exploration Guide Answers eBook Subscription Services
 - Covalent Bonding Exploration Guide Answers Budget-Friendly Options
6. Navigating Covalent Bonding Exploration Guide Answers eBook Formats
 - ePub, PDF, MOBI, and More
 - Covalent Bonding Exploration Guide Answers Compatibility with Devices
 - Covalent Bonding Exploration Guide Answers Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Covalent Bonding Exploration Guide Answers
 - Highlighting and Note-Taking Covalent Bonding Exploration Guide Answers
 - Interactive Elements Covalent Bonding Exploration Guide Answers
8. Staying Engaged with Covalent Bonding Exploration Guide Answers
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Covalent Bonding Exploration Guide Answers
9. Balancing eBooks and Physical Books Covalent Bonding Exploration Guide Answers
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Covalent Bonding Exploration Guide Answers
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Covalent Bonding Exploration Guide Answers
 - Setting Reading Goals Covalent Bonding Exploration Guide Answers
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Covalent Bonding Exploration Guide Answers
 - Fact-Checking eBook Content of Covalent Bonding Exploration Guide Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
- Exploring Educational eBooks

14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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