

CHAPTER 6 REVIEW*Chemical Bonding***MIXED REVIEW**

SHORT ANSWER Answer the following questions in the space provided.

1. Name the type of energy that is a measure of strength for each of the following types of bonds:

_____ lattice energy _____	a. ionic bond
_____ bond energy _____	b. covalent bond
_____ enthalpy of vaporization _____	c. metallic bond

2. Use the electronegativity values shown in **Figure 20**, on page 161 of the text, to determine whether each of the following bonds is nonpolar covalent, polar covalent, or ionic.

_____ ionic _____	a. H—F	_____ nonpolar covalent _____	d. H—H
_____ ionic _____	b. Na—Cl	_____ polar covalent _____	e. H—C
_____ polar covalent _____	c. H—O	_____ polar covalent _____	f. H—N

3. How is a hydrogen bond different from an ionic or covalent bond?

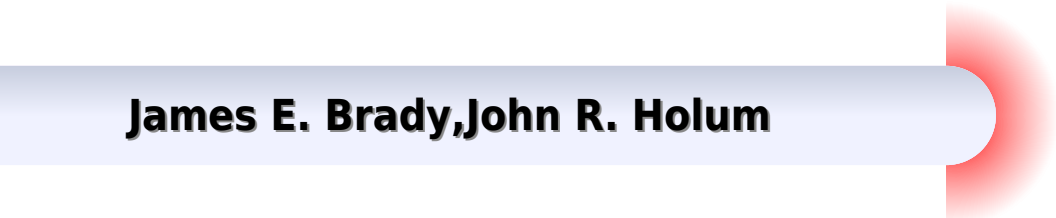
A hydrogen bond is a dipole-dipole attraction between a partially positive hydrogen
atom and the unshared electron pair of a strongly electronegative atom such as O,
N, or F. Unlike ionic or covalent bonds, in which electrons are given up or shared,
the hydrogen bond is a weaker attraction. Hydrogen bonds are generally
intermolecular, while ionic and covalent bonds occur between ions or atoms
respectively.

4. H₂S and H₂O have similar structures and their central atoms belong to the same group. Yet H₂S is a gas at room temperature and H₂O is a liquid. Use bonding principles to explain why this is.

Oxygen has higher electronegativity than sulfur, which creates a highly polar bond.
Increased polarity in H₂O bonds means a stronger intermolecular attraction, making
water a liquid at room temperature. Hydrogen bonding exists between water
molecules, but not between hydrogen sulfide molecules.

Covalent Bonding Guided And Study Workbook Answers

James E. Brady, John R. Holum



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1. Understanding the eBook Covalent Bonding Guided And Study Workbook Answers
 - The Rise of Digital Reading Covalent Bonding Guided And Study Workbook Answers
 - Advantages of eBooks Over Traditional Books
2. Identifying Covalent Bonding Guided And Study Workbook Answers
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Covalent Bonding Guided And Study Workbook Answers
 - User-Friendly Interface
4. Exploring eBook Recommendations from Covalent Bonding Guided And Study Workbook Answers
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 - Adjustable Fonts and Text Sizes of Covalent Bonding Guided And Study Workbook Answers
 - Highlighting and Note-Taking Covalent Bonding Guided And Study Workbook Answers
 - Interactive Elements Covalent Bonding Guided And Study Workbook Answers
8. Staying Engaged with Covalent Bonding Guided And Study Workbook Answers
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 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Covalent Bonding Guided And Study Workbook Answers
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 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Covalent Bonding Guided And Study Workbook Answers
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 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
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 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Covalent Bonding Guided And Study Workbook Answers
 - Fact-Checking eBook Content of Covalent Bonding Guided And Study Workbook Answers
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

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