

# Bond Energy

What makes a reaction endothermic or exothermic?

## Why?

Chemical reactions can be a lot like playing with Legos®—you must take apart part of your last creation before you can replace it with something new. For many chemical reactions, we have to first break bonds in the reactants before we can put the atoms back into a new arrangement to form the products. Both of these processes involve changes in energy. The net energy change for a reaction is called the **heat of reaction** or the **change in enthalpy** ( $\Delta H$ ). In this activity we will look at one way energy changes can be approximated for chemical reactions.

## Model 1 – Breaking and Forming Bonds

	Reaction	Change in Energy (kJ/mol <sub>rxn</sub> )		Reaction	Change in Energy (kJ/mol <sub>rxn</sub> )
A)	$\text{PCl}_3(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{Cl}(\text{g})$	+ 966.7 <i>broken</i>	E)	$\text{P}(\text{g}) + 3\text{Cl}(\text{g}) \rightarrow \text{PCl}_3(\text{g})$	- 966.7 <i>formed</i>
B)	$\text{PCl}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{Cl}(\text{g})$	+ 1,297.9 <i>broken</i>	F)	$\text{P}(\text{g}) + 5\text{Cl}(\text{g}) \rightarrow \text{PCl}_5(\text{g})$	- 1,297.9 <i>formed</i>
C)	$\text{PF}_3(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{F}(\text{g})$	+ 1,470.4 <i>broken</i>	G)	$\text{P}(\text{g}) + 3\text{F}(\text{g}) \rightarrow \text{PF}_3(\text{g})$	- 1,470.4 <i>formed</i>
D)	$\text{PF}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{F}(\text{g})$	+ 2,305.4 <i>broken</i>	H)	$\text{P}(\text{g}) + 5\text{F}(\text{g}) \rightarrow \text{PF}_5(\text{g})$	- 2,305.4 <i>formed</i>

- Identify four reactions in Model 1 where bonds are being broken. Write "broken" in the Change in Energy box in the table for each reaction you identify.
- Identify four reactions in Model 1 where bonds are being formed. Write "formed" in the Change in Energy box in the table for each reaction you identify.
- Circle the correct word to complete each sentence below.
  - When bonds are (broken/formed) there is a positive energy change.
  - Breaking bonds is (endothermic/exothermic).
  - When bonds are (broken/formed) there is a negative energy change.
  - Forming bonds is (endothermic/exothermic).
- Find two reactions in Model 1 that are exact opposites of each other, that is, one reaction is the reverse of the other reaction.
  - How do the changes in energy for the reverse reactions compare?
 

*The changes in energy for these reactions are the same, but opposite in sign.*
  - Explain your answer to part a considering what you learned from Questions 1–3 about bond breaking and bond formation.
 

*If a given amount of energy is required to break a bond, then the same amount of energy will be released when the bond is formed.*

# Bond Energy Graph Pogil

**Lorcan Stuart Peter Stillwell Smith**

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the author's name.

## **Bond Energy Graph Pogil:**

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*Introduction to Bond Graphs and their Applications* Jean U. Thoma, 2016-05-04 Introduction to Bond Graphs and Their Applications is an introductory text on bond graphs and their applications in the field of engineering The applications of bond graphs in mechanical engineering and design fluid mechanics electronic data processing and thermal and thermodynamic systems are discussed This book is comprised of eight chapters and begins by comparing the different kinds of graphs diagrams and models before turning to the fundamentals of bond graphs The next chapter introduces the reader to the systematic application of bond graphs in mechanical engineering and design fluid power engineering sometimes called oil hydraulics electrotechnique and electronics and thermodynamics The use of bond graphs in automatic computer programming with the ENPORT program is also described The final chapter is devoted to inertia and resistance fields linear two ports in different causalities thermodynamics of flow processes electromechanical components systems with distributed parameters and force and velocity as effort or flow This monograph is intended primarily for all engineers interested in representing simple or complex engineering systems and should also be of value to students in the different engineering disciplines mechanics fluid mechanics and electronics with electromechanical power conversion or thermodynamics

**Chemical Bonds and Bonds Energy** R Sanderson, 2012-12-02 Chemical Bonds and Bonds Energy Second Edition provides information pertinent to the fundamental aspects of contributing bond energy and bond dissociation energy This book explores the values that are useful in the interpretation of significant phenomena such as product distribution and reaction mechanisms Organized into 12 chapters this edition begins with an overview of the quantitative relationship among

three basic properties of an atom namely nonpolar covalent radius electronegativity and homonuclear single covalent bond energy This text then examines the quantitative means of evaluating the partial atomic charges that result from initial differences in the electromagnetivity of atoms that form a compound Other chapters consider the recognition of the reduction of bond weakening not by multiplicity and in certain types of single covalent bonds The final chapter deals with the application of the principal ideas and techniques to the oxidation of ethane This book is a valuable resource for organic and inorganic chemists

Chemical Bonds and Bonds Energy Robert Thomas Sanderson,1976-06-28 Chemical Bonds and Bonds Energy Second Edition provides information pertinent to the fundamental aspects of contributing bond energy and bond dissociation energy This book explores the values that are useful in the interpretation of significant phenomena such as product distribution and reaction mechanisms Organized into 12 chapters this edition begins with an overview of the quantitative relationship among three basic properties of an atom namely nonpolar covalent radius electronegativity and homonuclear single covalent bond energy This text then examines the quantitative means of evaluating the partial atomic charges that result from initial differences in the electromagnetivity of atoms that form a compound Other chapters consider the recognition of the reduction of bond weakening not by multiplicity and in certain types of single covalent bonds The final chapter deals with the application of the principal ideas and techniques to the oxidation of ethane This book is a valuable resource for organic and inorganic chemists

**Phase Co-ordinate Bond Graph Analysis of Multiport Energy Systems**,1974 **Energy Bond Graphs** Peder Voetmann Christiansen,2005 **Phase co-ordinate bond graph analysis of multiport energy systems** Mohammed Ridzuan Bin Haji Mohammed Salleh,1974 **Bond Graph Modelling of Exergy in Integrated Energy Systems** Robin Wardle,2022 Phase Coordinate Bond Graph Analysis of Multiport Energy Systems Mohammed Ridzuan Salleh,1974 **Comprehensive Handbook of Chemical Bond Energies** Yu-Ran Luo,2007-03-09

Understanding the energy it takes to build or break chemical bonds is essential for scientists and engineers in a wide range of innovative fields including catalysis nanomaterials bioengineering environmental chemistry and space science Reflecting the frequent additions and updates of bond dissociation energy BDE data throughout the literature the Comprehensive Handbook of Chemical Bond Energies compiles the most recent experimental BDE data for more than 19 600 bonds of 102 elements The author organizes the data by bond type functional group bond order bond degree molecular size and structure for ease of use Data can also be located using the Periodic table The book presents data for organic molecules biochemicals and radicals as well as clusters ions hydrogen and surface bonded species van der Waals complexes isotopic species and halogen clusters complexes It also introduces entirely new data for inorganics and organometallics The final chapter summarizes the heats of formation for atoms inorganic organic radicals and monoatomic ions in the gas phase The Comprehensive Handbook of Chemical Bond Energies offers quick access to experimental BDE data in the most inclusive well organized and up to date collection available today

**Energy Bond Graphs** ,2003 *The Chemical Bond in Inorganic*

*Chemistry* Ian David Brown, 2006 The bond valence model is a recently developed model of the chemical bond in inorganic chemistry that complements the bond model widely used in organic chemistry It is simple quantitative intuitive and predictive no more than a pocket calculator is needed to calculate it This book focuses on the theory that underlies the model and shows how it has been used in physics materials science chemistry mineralogy soil science and molecular biology

*Energy Bond Graphs* Peder Voetmann Christiansen, Roskilde Universitetscenter. Institut for Studiet af Matematik og Fysik samt deres Funktioner i Undervisning, Forskning og Anvendelser, 2005 *Bond Graph Techniques for Dynamic Systems in Engineering and Biology* Dean Karnopp, 1979 *Bond Graphs for Causal Explanations* J. L. Top, 1992 **Energy**, 1976 **A Pictorial Approach to Molecular Bonding** John G. Verkade, 2012-12-06 With the development of accurate molecular calculations in recent years useful predictions of molecular electronic properties are currently being made It is therefore becoming increasingly important for the non theoretically oriented chemist to appreciate the underlying principles governing molecular orbital formation and to distinguish them from the quantitative details as associated with particular molecules It seems highly desirable then that the non theoretician be able to deduce results of general validity without esoteric mathematics In this context pictorial reasoning is particularly useful Such an approach is virtually indispensable if bonding concepts are to be taught to chemistry students early in their careers Undergraduate chemistry majors typically find it difficult to formulate molecular orbital schemes especially delocalized ones for molecules more complicated than diatomics The major reason for this regrettable situation is the general impracticability of teaching group theory before students take organic and inorganic courses wherein the applications of these concepts are most beneficial Consequently many students graduate with the misconception that the ground rules governing bonding in molecules such as  $\text{NH}_3$  are somehow different from those which apply to aromatic systems such as  $\text{C}_6\text{H}_6$  Conversely seniors and many graduate students are usually only vaguely if at all aware that sigma bonding like extended pi bonding can profitably be described in a delocalized manner when discussing the UV photoelectron spectrum of CH for example *Introduction to Bond Graphs and Their Applications* Jean Ulrich Thoma, 1975 *Bond Graph Modelling of Physical Systems* Lorcan Stuart Peter Stillwell Smith, 1993 *Spatial-energy Principles of the Processes for Complex Structure Formation* G. A. Korabely, 2005

## Unveiling the Power of Verbal Beauty: An Emotional Sojourn through **Bond Energy Graph Pogil**

In some sort of inundated with screens and the cacophony of fast conversation, the profound power and mental resonance of verbal art often fade into obscurity, eclipsed by the regular barrage of noise and distractions. However, nestled within the musical pages of **Bond Energy Graph Pogil**, a fascinating work of fictional brilliance that impulses with fresh feelings, lies an wonderful trip waiting to be embarked upon. Published with a virtuoso wordsmith, that exciting opus guides viewers on an emotional odyssey, softly revealing the latent potential and profound affect embedded within the complicated web of language. Within the heart-wrenching expanse of the evocative examination, we will embark upon an introspective exploration of the book is main themes, dissect its fascinating publishing type, and immerse ourselves in the indelible impact it leaves upon the depths of readers souls.

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