

**Bond Enthalpy (Bond Energy)**

**Bond enthalpy**, also known as **bond energy**, is defined as the amount of energy required to break a bond. These values are positive, indicating that bond breaking requires energy and is endothermic. Bond enthalpies are reported in kilojoules per mol (kJ/mol). The average energy for breaking a hydrogen-hydrogen bond is 436 kJ/mol.

In a chemical reaction several bonds are broken and formed. For example, in the reaction below a hydrogen-hydrogen bond is broken and a fluorine-fluorine bond is broken. Two hydrogen-fluorine bonds are formed. The overall enthalpy change for this reaction is calculated below:



Bond Energy: 436 kJ/mol



Bond Energy: 155 kJ/mol

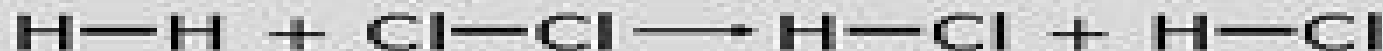


Bond Energy: 567 kJ/mol

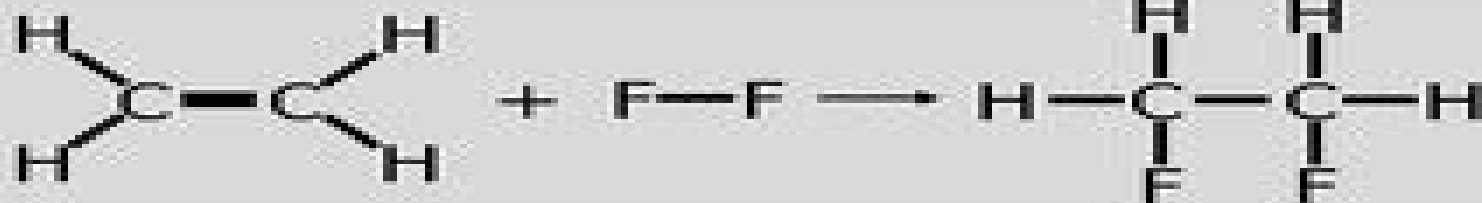


$$\begin{aligned}\Delta H_{\text{rxn}} &= [\text{energy used for breaking bonds}] - [\text{energy formed in making bonds}] \\ &= [436 \text{ kJ/mol} + 155 \text{ kJ/mol}] - [2(567 \text{ kJ/mol})] \\ &= -543 \text{ kJ/mol}\end{aligned}$$

Estimate the enthalpy of reaction ( $\Delta H_{\text{rxn}}$ ) using bond energy values.



Work area



Work area

# Bond Energy Key

**David M. Hillis, Mary V. Price, Richard  
W. Hill, David W. Hall, Marta J.  
Laskowski**

## **Bond Energy Key:**

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