

## Chapter 6: Chemical Bonding Guided Notes

### I. Why do Atoms Combine?

#### A. Chemical Formula

1. The chemical formula shows:
  - a. The \_\_\_\_\_ in the compound.
  - b. The ratio of their \_\_\_\_\_

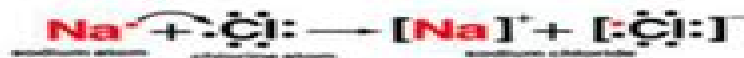


#### B. Chemical Bond

1. Strong attractive force between \_\_\_\_\_ or ions in a molecule or compound.
2. Formed by:
  - a. \_\_\_\_\_ electrons, (e-) by losing or gaining electrons.
  - b. \_\_\_\_\_ electrons (e-)

#### C. Stability

1. Octet Rule
  - a. Most atoms form bonds in order to have \_\_\_\_\_ valence electrons (e-).
  - b. \_\_\_\_\_ outer energy level.
  - c. Want to be like the Noble Gases (group 18).
2. Stability is the driving force behind bond formation. All atoms bond to become stable.
3. Transferring e-



4. Sharing e-



### II. Kinds of Chemical Bonds

#### A. Ionic Bond

1. Attraction between 2 \_\_\_\_\_ charged ions.
2. Ions are \_\_\_\_\_ atoms.
3. A \_\_\_\_\_ is a \_\_\_\_\_ charged ion. Forms when an atom loses an electron.  
Formed by \_\_\_\_\_.
4. A \_\_\_\_\_ is a \_\_\_\_\_ charged ion. Forms when an atom gains an electron.  
Formed by \_\_\_\_\_.
5. Formed by \_\_\_\_\_ electrons (e-) from a \_\_\_\_\_ to a \_\_\_\_\_.
6. Ions form a 3-D crystal lattice.



# Bonding Guided Notes Answers

**Johnny P. Bradley**



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