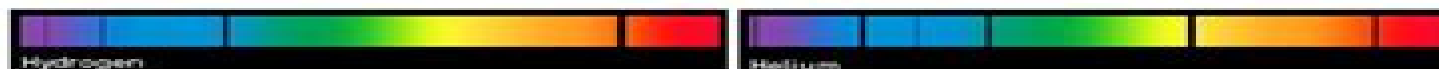


## Student Exploration: Bohr Model: Introduction

When light passes through a gas, certain wavelengths of the light are absorbed. The result is a unique **absorption spectrum**. Two examples are shown below.



1. What colors of light are absorbed by hydrogen gas?  
violet, blue, green, and red
2. What colors of light are absorbed by helium gas?  
violet, blue, green, yellow, and red

### Gizmo Warm-up

In 1913, Niels Bohr proposed that the unique spectral lines created by different elements were related to the way electrons were arranged around the nucleus. The *Bohr Model: Introduction* Gizmo explores this connection.

The **laser** shown in the Gizmo can emit **photons**, or particles of light, at a variety of wavelengths. The energy of a photon, measured in **electron volts (eV)**, is inversely proportional to its wavelength. Photons that pass through the gas are detected by the photon detector at the right.



1. With the **Energy (eV)** set to 1 eV, click **Fire**. Did the photon go straight through the gas in the tube, or was it absorbed by the gas?  
the photon went straight through the gas tube.
2. Set the **Energy (eV)** to 4 eV, and click **Fire**. What happened this time?  
the photon was absorbed by the gas.

### Activity A: Absorption spectra

#### Get the Gizmo ready:

- On the **SIMULATION** pane, select **Lamp**.
- Check that **Gas C** is selected.



**Introduction:** The smaller the wavelength of a photon, the greater its energy. Humans can see photons with wavelengths between 700 nanometers (red) and 400 nanometers (violet), corresponding to energies between 1.8 and 3.1 eV.

**Question:** What does the absorption spectrum of an element indicate about its electron configuration?

# Bohr Model Of Energy Gizmo Answers

**K Payea**



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## **Table of Contents Bohr Model Of Energy Gizmo Answers**

1. Understanding the eBook Bohr Model Of Energy Gizmo Answers
  - The Rise of Digital Reading Bohr Model Of Energy Gizmo Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Bohr Model Of Energy Gizmo Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Bohr Model Of Energy Gizmo Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Bohr Model Of Energy Gizmo Answers
  - Personalized Recommendations
  - Bohr Model Of Energy Gizmo Answers User Reviews and Ratings

- Bohr Model Of Energy Gizmo Answers and Bestseller Lists
- 5. Accessing Bohr Model Of Energy Gizmo Answers Free and Paid eBooks
  - Bohr Model Of Energy Gizmo Answers Public Domain eBooks
  - Bohr Model Of Energy Gizmo Answers eBook Subscription Services
  - Bohr Model Of Energy Gizmo Answers Budget-Friendly Options
- 6. Navigating Bohr Model Of Energy Gizmo Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Bohr Model Of Energy Gizmo Answers Compatibility with Devices
  - Bohr Model Of Energy Gizmo Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Bohr Model Of Energy Gizmo Answers
  - Highlighting and Note-Taking Bohr Model Of Energy Gizmo Answers
  - Interactive Elements Bohr Model Of Energy Gizmo Answers
- 8. Staying Engaged with Bohr Model Of Energy Gizmo Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Bohr Model Of Energy Gizmo Answers
- 9. Balancing eBooks and Physical Books Bohr Model Of Energy Gizmo Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Bohr Model Of Energy Gizmo Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Bohr Model Of Energy Gizmo Answers
  - Setting Reading Goals Bohr Model Of Energy Gizmo Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Bohr Model Of Energy Gizmo Answers
  - Fact-Checking eBook Content of Bohr Model Of Energy Gizmo Answers
  - Distinguishing Credible Sources

13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
  - Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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