

Bond Energy

What makes a reaction endothermic or exothermic?

Why?

Chemical reactions can be a lot like playing with Legos®—you must take apart part of your last creation before you can replace it with something new. For many chemical reactions, we have to first break bonds in the reactants before we can put the atoms back into a new arrangement to form the products. Both of these processes involve changes in energy. The net energy change for a reaction is called the **heat of reaction** or the **change in enthalpy** (ΔH). In this activity we will look at one way energy changes can be approximated for chemical reactions.

Model 1 – Breaking and Forming Bonds

	Reaction	Change in Energy (kJ/mol _{rxn})		Reaction	Change in Energy (kJ/mol _{rxn})
A)	$\text{PCl}_3(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{Cl}(\text{g})$	+ 966.7 <i>broken</i>	E)	$\text{P}(\text{g}) + 3\text{Cl}(\text{g}) \rightarrow \text{PCl}_3(\text{g})$	- 966.7 <i>formed</i>
B)	$\text{PCl}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{Cl}(\text{g})$	+ 1,297.9 <i>broken</i>	F)	$\text{P}(\text{g}) + 5\text{Cl}(\text{g}) \rightarrow \text{PCl}_5(\text{g})$	- 1,297.9 <i>formed</i>
C)	$\text{PF}_3(\text{g}) \rightarrow \text{P}(\text{g}) + 3\text{F}(\text{g})$	+ 1,470.4 <i>broken</i>	G)	$\text{P}(\text{g}) + 3\text{F}(\text{g}) \rightarrow \text{PF}_3(\text{g})$	- 1,470.4 <i>formed</i>
D)	$\text{PF}_5(\text{g}) \rightarrow \text{P}(\text{g}) + 5\text{F}(\text{g})$	+ 2,305.4 <i>broken</i>	H)	$\text{P}(\text{g}) + 5\text{F}(\text{g}) \rightarrow \text{PF}_5(\text{g})$	- 2,305.4 <i>formed</i>

- Identify four reactions in Model 1 where bonds are being broken. Write "broken" in the Change in Energy box in the table for each reaction you identify.
- Identify four reactions in Model 1 where bonds are being formed. Write "formed" in the Change in Energy box in the table for each reaction you identify.
- Circle the correct word to complete each sentence below.
 - When bonds are (broken/formed) there is a positive energy change.
 - Breaking bonds is (endothermic/exothermic).
 - When bonds are (broken/formed) there is a negative energy change.
 - Forming bonds is (endothermic/exothermic).
- Find two reactions in Model 1 that are exact opposites of each other, that is, one reaction is the reverse of the other reaction.

- How do the changes in energy for the reverse reactions compare?

The changes in energy for these reactions are the same, but opposite in sign.

- Explain your answer to part a considering what you learned from Questions 1–3 about bond breaking and bond formation.

If a given amount of energy is required to break a bond, then the same amount of energy will be released when the bond is formed.

Bond Energy Pogil Extension

**Christopher S. Bond, Missouri
Transition Government**



Bond Energy Pogil Extension:

Comprehensive Handbook of Chemical Bond Energies Yu-Ran Luo, 2007-03-09 Understanding the energy it takes to build or break chemical bonds is essential for scientists and engineers in a wide range of innovative fields including catalysis nanomaterials bioengineering environmental chemistry and space science Reflecting the frequent additions and updates of bond dissociation energy BDE data throughout the literature the Comprehensive Handbook of Chemical Bond Energies compiles the most recent experimental BDE data for more than 19 600 bonds of 102 elements The author organizes the data by bond type functional group bond order bond degree molecular size and structure for ease of use Data can also be located using the Periodic table The book presents data for organic molecules biochemicals and radicals as well as clusters ions hydrogen and surface bonded species van der Waals complexes isotopic species and halogen clusters complexes It also introduces entirely new data for inorganics and organometallics The final chapter summarizes the heats of formation for atoms inorganic organic radicals and monoatomic ions in the gas phase The Comprehensive Handbook of Chemical Bond Energies offers quick access to experimental BDE data in the most inclusive well organized and up to date collection available today **CHEMICAL BONDS AND BOND ENERGY.** RT. SANDERSON, 1971 **Bond Orders and Energy Components** , 2013 **Bond Orders and Energy Components** István Mayer, 2016-10-14 While modern computational methods can provide us with the wave function of a molecule in numerical form most computer programs lack the sophisticated tools needed to extract chemical concepts from these wave functions Saving researchers vast time and potential confusion this volume collects and organizes those validated tools currently scattered throughout the literature and details their application It provides immediate access for those needing to calculate such critical factors as bond order and valence indices and atomic and diatomic contributions to molecular energy Supporting material is available for download from the authors continually updated website *Empirical Bond Energies* FRANK JR. WILSON, AIR FORCE FLIGHT TEST CENTER EDWARDS AFB CALIF., Air Force Flight Test Center (U.S.), 1961 A consistent set of bond energy values are calculated on the assumption that bond energy nonconstancies are a minimum when determined from compounds in which all the constituent elements exhibit their normal valency Several empirical rules for estimating the heats of vaporization and fusion are included The report contains sufficient data for the estimation of heats of formation of compounds In many cases estimations can be made within a few kilocalories Author **Municipal Bond Financing for Renewable Energy Systems** Alice Levine, 1981 **Bond-based Corrections to Improve Atomization Energy Calculations** Neelabh Maheshwari, 2012 **Energy Refinements to Glass Structure** Patrick Joseph Hood, 1985 Kit Bond's Position Paper No. 4 Christopher S. Bond, Missouri Transition Government, 1980 **S.C.F. Calculations of Bond Properties and Potential Energy Surfaces** , 1971 **Main Group Chemistry** , 2013 **Phase Co-ordinate Bond Graph Analysis of Multiport Energy Systems** , 1974 *The Basics of Bond Financings of the State Environmental Improvement and Energy Resources Authority* Angela

Fick Braly, Beverly A. Marcin, Lewis, Rice & Fingersh, Missouri. Department of Natural Resources, 1994 **Utilizing the Bond Order-bond Length Relationship for Reorganization Energy Prediction of Polyaromatic Hydrocarbons** □□ □, 2023 **Municipal Bond Financing for Renewable Energy Systems** Alice Levine, 1981 Phase co-ordinate bond graph analysis of multiport energy systems Mohammed Ridzuan Bin Haji Mohammed Salleh, 1974 Bond Selective Chemistry Beyond the Adiabatic Approximation , 1993 The adiabatic Born Oppenheimer potential energy surface approximation is not valid for reaction of a wide variety of energetic materials and organic fuels coupling between electronic states of reacting species plays a key role in determining the selectivity of the chemical reactions induced This research program initially studies this coupling in 1 selective C Br bond fission in 1 3 bromiodopropane 2 C S S H bond fission branching in CH sub 3 SH and 3 competition between bond fission channels and H sub 2 elimination in CH sub 3 NH sub 2

Energy and Bonding Michael Hudson (chimiste), 1969 *Energy Bond Graphs* Peder Voetmann Christiansen, Roskilde Universitetscenter. Institut for Studiet af Matematik og Fysik samt deres Funktioner i Undervisning, Forskning og Anvendelser, 2005

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Table of Contents Bond Energy Pogil Extension

1. Understanding the eBook Bond Energy Pogil Extension
 - The Rise of Digital Reading Bond Energy Pogil Extension
 - Advantages of eBooks Over Traditional Books
2. Identifying Bond Energy Pogil Extension
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Bond Energy Pogil Extension
 - User-Friendly Interface
4. Exploring eBook Recommendations from Bond Energy Pogil Extension
 - Personalized Recommendations
 - Bond Energy Pogil Extension User Reviews and Ratings

- Bond Energy Pogil Extension and Bestseller Lists
- 5. Accessing Bond Energy Pogil Extension Free and Paid eBooks
 - Bond Energy Pogil Extension Public Domain eBooks
 - Bond Energy Pogil Extension eBook Subscription Services
 - Bond Energy Pogil Extension Budget-Friendly Options
- 6. Navigating Bond Energy Pogil Extension eBook Formats
 - ePub, PDF, MOBI, and More
 - Bond Energy Pogil Extension Compatibility with Devices
 - Bond Energy Pogil Extension Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Bond Energy Pogil Extension
 - Highlighting and Note-Taking Bond Energy Pogil Extension
 - Interactive Elements Bond Energy Pogil Extension
- 8. Staying Engaged with Bond Energy Pogil Extension
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Bond Energy Pogil Extension
- 9. Balancing eBooks and Physical Books Bond Energy Pogil Extension
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Bond Energy Pogil Extension
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Bond Energy Pogil Extension
 - Setting Reading Goals Bond Energy Pogil Extension
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Bond Energy Pogil Extension
 - Fact-Checking eBook Content of Bond Energy Pogil Extension
 - Distinguishing Credible Sources

-
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
 14. Embracing eBook Trends
 - Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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