

Review Unit 6 – Bonding

I. Bond Types: Fill in the table to show the basics of bond types.

Bond	Type of atoms involved	Electrons transferred or shared?	Type of Compound formed	What holds the bond together?
Ionic	M/N	Trans	Ionic Polar	e- transfer, electrostatic forces
Covalent	N/N	Shared	Covalent Polar and Nonpolar	Sharing of e-
Metallic	M	Shared	Alloys	Sharing of e-

II. Short Answer:

- Why are some covalent bonds polar? **Higher electronegative atoms unequally share the electrons in the bond.**
- Which of the following compounds express polarity? To which element are the compounds polar too?
 - H_2O **Yes O**
 - HF **Yes F**
 - CCl_4 **No**
 - CO_2 **Yes O**
- What are Binary Acids? List 4 and include the formula for each. **Binary acids consist of H and another element. HF HCl HI HBr**
- What are Tertiary Acids? List 4 and include the formula for each. **Tertiary Acids include H and a polyatomic ion in most cases. HNO_3 , HClO_3 , HPO_4 , and H_2SO_3**
- How does electronegativity relate to the strength of an acid? **The higher the electronegativity the greater the strength of the acid.**
- Is it possible for a bond to be partially ionic? Explain. **No, only whole electrons are transferred in an ionic bond resulting in a definite charge.**
- What is an ion? How does it form? **An ion is a charged atom that forms from the gain or loss of electrons.**
- What is a polyatomic ion? What kind of bond holds the atoms in a polyatomic ion together? **A polyatomic ion is a charged, covalently bonded molecule that combines with a metal to form an ionic compound.**
- Show the electron configuration for each of the atoms below and the configuration for the ion as well.
 - Ca
 $1s^2 2s^2 2p^6 3s^2 3p^4 4s^2$
 Ion: $1s^2 2s^2 2p^6 3s^2 3p^4$
 - N
 $1s^2 2s^2 2p^3$
 Ion: $1s^2 2s^2 2p^6$
 - Hg (ion: 2+)
 $[\text{Xe}] 6s^2 4f^{14} 5d^{10}$

Bonding Basics Review Answer

EW Minium



Bonding Basics Review Answer:

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