

CHEM 101 Quiz #6 Chapter 6 ANSWERS**1. (4 points)**

The specific heat (s) of iron (Fe) is $0.40 \text{ J/(g}\cdot\text{K)}$. Calculate the heat effect q when 20 g of iron is heated from 100°C to 200°C . Include the sign!

$$q = m \cdot s \cdot \Delta T \text{ for } \Delta T \text{ in } ^\circ\text{C} \quad q = 20 \times 0.40 \times (200 - 100) = +800 \text{ J}$$

You must include the units and sign in the answer!

2. (4 points)

The Molar Mass of Mg = 24.0 g/mol .

Calculate the enthalpy change when 2.40 g Mg(s) reacts with excess oxygen. Include the sign!

$$2.40 \text{ g Mg} = 2.40/24.0 = 0.1 \text{ mol} \rightarrow \Delta H = 0.1(\text{mol}) \times (-600)(\text{kJ/mol}) = -60 \text{ kJ}$$

Again, your answer must include both the sign and the units.

3. (2 points) Which one of the following four reactions is a *formation reaction*? Circle the correct answer.



Look up the definition of a formation reaction in Chang or in your lecture notes: the compound is formed from its elements, and the elements must be as they naturally occur so e.g. Cl_2 not Cl , and H_2 not H !

CHEM 101 Section L52 Quiz #6b**1. (4 points)**

The specific heat (s) of copper (Cu) is $0.38 \text{ J/(g}\cdot\text{K)}$. Calculate the heat effect q when 10 g of copper is heated from 100°C to 300°C . Include the sign!

$$q = m \cdot s \cdot \Delta T \text{ for } \Delta T \text{ in } ^\circ\text{C} \quad q = 10 \times 0.38 \times (300 - 100) = +760 \text{ J}$$

You must include the units and sign in the answer!

2. (4 points)

The Molar Mass of Ca = 40.0 g/mol .

Calculate the enthalpy change ΔH when 8.0 g Ca(s) reacts with excess oxygen. Include the sign!!

$$8.0 \text{ g Ca} = 8.0/40 = 0.2 \text{ mol} \rightarrow \Delta H = 0.2(\text{mol}) \times (-650)(\text{kJ/mol}) = -130 \text{ kJ}$$

Again, your answer must include both the sign and the units.

3. (2 points) Which one of the following reactions is a *formation reaction*? Circle the correct answer.



Chapter 6 Chemistry Test

G. Philip



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Laboratory and Diagnostic Testing in Ambulatory Care E-Book Martha (Marti) Garrels, Carol S. Oatis, 2018-08-07
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outside of the physician's office explains what happens to the samples you send out. There's also information on CLIA and other government regulations and how they affect each procedure.

Laboratory Testing for Ambulatory Settings - E-Book Martha (Marti) Garrels, 2014-06-16. Learn the lab testing skills you need to know. *Laboratory Testing for Ambulatory Settings: A Guide for Health Care Professionals*, 2nd Edition, provides in-depth coverage of the most common procedures and techniques of all the new CLIA waived point-of-care tests along with some moderately complex tests. Clear step-by-step instructions and full-color photographs make it easy to master each test and procedure. Written by noted educators Marti Garrels and Carol S. Oatis, this edition adds a new chapter on toxicology and information on five new procedures. A companion Evolve website lets you practice clinical laboratory skills. Complete coverage includes the most common CLIA waived tests for any healthcare professional in the ambulatory setting. A triad organization gives chapters a consistent, easy-to-follow format: Fundamental Concepts (basic information related to tests and procedures), CLIA Waived Procedures (step-by-step instructions for CLIA waived tests), and Advanced Concepts (further application of basic knowledge and skills towards a higher level of critical thinking and decision making, such as handling non-CLIA waived tests). Procedure boxes provide clear step-by-step instructions along with numerous full-color photos and illustrations. Key terms are defined and reinforced within each chapter. Common abbreviations associated with CLIA waived testing are cited at the beginning of chapters. 710 review questions conclude each chapter to reinforce learning. A companion Evolve website includes various activities and exercises to enhance learning with problem-solving scenarios. A workbook matches the chapters in the textbook, offering activities and exercises to reinforce laboratory concepts, terminology, and procedures. Skills sheets help you work through the competency-based procedures and meet government standards for good laboratory practice. Sold separately. Over 60 new photographs and drawings clarify topics and show examples of laboratory specimens so you will be able to identify them on the job. New chapter on toxicology. More extensive coverage of working with the microscope prepares you for the lab. Updated, expanded information about quality control and quality assurance provides relevant information so you can accurately and effectively perform in the lab. Updates on proper collection and processing of urine specimens, microbiology specimens, blood, capillary, and venipuncture specimens, which includes new urine culture using vacutainer system, new drawings for urine and influenza specimen collections, and the latest order of blood draw including the plasma separator tube, ensure that you have the most current information. Five new procedures with corresponding skill check-off sheets help you understand the most up-to-date protocols. Clinitek Analyzer, Standard Hematocrit, INRatio, New A1c, Ki iFOB method for fecal occult blood. *Phlebotomy Exam Review, Enhanced Edition*, Ruth E. McCall, 2020-06-19. *Phlebotomy Exam Review, Enhanced*, Seventh Edition, thoroughly prepares students for any of the national certification exams in phlebotomy.

[5 Steps to a 5 AP Chemistry, 2014-2015 Edition](#) Richard H. Langley, John Moore, 2013-08-02. A PERFECT PLAN for the PERFECT SCORE STEP 1. Set up your study plan with three customized study schedules. STEP 2: Determine your readiness with an AP-style diagnostic exam. STEP 3:

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