

## Chapter 5—Electrons in Atoms

---

### MATCHING

Match each item with the correct statement below.

- |                          |                                    |
|--------------------------|------------------------------------|
| a atomic orbital         | d ground state                     |
| b aufbau principle       | e Pauli exclusion principle        |
| c electron configuration | f Heisenberg uncertainty principle |
- 
1. region of high probability of finding an electron **A**
  2. states the impossibility of knowing both velocity and position of a moving particle at the same time **F**
  3. lowest energy level **D**
  4. tendency of electrons to enter orbitals of lowest energy first **B**
  5. arrangement of electrons around atomic nucleus **C**
  6. each orbital has at most two electrons **e**

### MULTIPLE CHOICE

1. In Bohr's model of the atom, where are the electrons and protons located?
  - a The electrons move around the protons, which are at the center of the atom.
  - b The electrons and protons move throughout the atom.
  - c The electrons occupy fixed positions around the protons, which are at the center of the atom.
  - d The electrons and protons are located throughout the atom, but they are not free to move.

ANS: A

PTS: 1

DIF: L2

REF: p. 128

OBJ: 5.1.2

2. In the Bohr model of the atom, an electron in an orbit has a fixed \_\_\_\_\_.
  - a Position
  - b Color
  - c energy
  - d size

ANS: C

PTS: 1

DIF: L2

REF: p. 128

OBJ: 5.1.2

3. How does the energy of an electron change when the electron moves closer to the nucleus?
  - a It decreases.
  - b It increases.
  - c It stays the same.
  - d It doubles.

ANS: A

PTS: 1

DIF: L2

REF: p. 128

OBJ: 5.1.3

4. The principal quantum number indicates what property of an electron?

# Chapter 5 Assessment Electrons In Atoms Answer Key

**A Gutmann**



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