

Chapter 5: Thermochemistry

Problem Set with Answers

1. A system receives 425 J of heat and delivers 425 J of work to its surroundings. What is the change in internal energy of the system (in J)?

Answer: 0 J

2. Complete combustion of 1.0 metric ton of coal (assuming pure carbon) to gaseous carbon dioxide releases 3.3×10^{10} J of heat. Convert this energy to (a) kilojoules; (b) kilocalories

Answer: (a) 3.3×10^7 kJ (b) 7.9×10^6 kcal

3. Classify the following processes as exothermic or endothermic: (a) freezing of water; (b) boiling of water; (c) digestion of food; (d) a person running; (e) a person growing; (f) wood being chopped; (g) heating with a furnace.

Answer: (a) exothermic (b) endothermic (c) exothermic (d) exothermic
(e) endothermic (f) endothermic (g) exothermic

4. Calculate q when 12.0 g of water is heated from 20 °C to 100 °C.

Answer: 4.0×10^3 J

5. A 295-g aluminum engine part at an initial temperature of 3.00°C absorbs 85.0 kJ of heat. What is the final temperature of the part (c of Al = 0.900 J/g·K)?

Answer: 323°C

6. Ethylene (C_2H_4) is the starting material for the preparation of polyethylene. Although typically made during the processing of petroleum, ethylene occurs naturally as a fruit-ripening hormone and as a component of natural gas.
(a) The heat of reaction for the combustion of C_2H_4 is -1411 kJ/mole. Write a balanced thermochemical equation for the combustion of C_2H_4 .
(b) How many grams of C_2H_4 must burn to give 70.0 kJ of heat?

Answer: (a) $\text{C}_2\text{H}_4 (\text{g}) + 3\text{O}_2 (\text{g}) \rightarrow 2\text{CO}_2 (\text{g}) + 2\text{H}_2\text{O} (\text{l})$; $\Delta H_{\text{rxn}} = -1411 \text{ kJ}$
(b) 1.39 g C_2H_4

Chapter 5 Thermochemistry Answers

N Noddings



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Understanding General Chemistry Atef Korchef, 2022-03-07 Understanding General Chemistry details the fundamentals of general chemistry through a wide range of topics relating the structure of atoms and molecules to the properties of matter. Written in an easy to understand format with helpful pedagogy to fuel learning, the book features main objectives at the beginning of each chapter, get smart sections, and a check your reading section at the end of each chapter. The text is filled with examples and practices that illustrate the concepts at hand. In addition, a summary and extensive MCQs, exercises, and problems with the corresponding answers and explanations are readily available. Additional features include Alerts, students to common mistakes and explains in simple ways and clear applications how to avoid these mistakes. Offers answers and comments alongside sample problems enabling students to self-evaluate their skill level. Includes powerful methods, easy steps, simple and accurate interpretations, and engaging applications to help students understand complex principles. Provides a bridge to more complex topics such as solid state chemistry, organometallic chemistry, chemistry of main group elements, inorganic chemistry, and physical chemistry. This introductory textbook is ideal for chemistry courses for non-science majors as well as health sciences and preparatory engineering students.

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