

Chemical Bonding

SECTION 3

SHORT ANSWER Answer the following questions in the space provided.

- A** The notation for sodium chloride, NaCl, stands for one
 (a) formula unit. (c) crystal.
 (b) molecule. (d) atom.
- D** In a crystal of an ionic compound, each cation is surrounded by a number of
 (a) molecules. (c) dipoles.
 (b) positive ions. (d) negative ions.
- B** Compared with the neutral atoms involved in the formation of an ionic compound, the crystal lattice that results is
 (a) higher in potential energy. (c) equal in potential energy.
 (b) lower in potential energy. (d) unstable.
- B** The lattice energy of compound A is greater in magnitude than that of compound B. What can be concluded from this fact?
 (a) Compound A is not an ionic compound.
 (b) It will be more difficult to break the bonds in compound A than those in compound B.
 (c) Compound B has larger crystals than compound A.
 (d) Compound A has larger crystals than compound B.
- B** The forces of attraction between molecules in a molecular compound are generally
 (a) stronger than the attractive forces among formula units in ionic bonding.
 (b) weaker than the attractive forces among formula units in ionic bonding.
 (c) approximately equal to the attractive forces among formula units in ionic bonding.
 (d) equal to zero.
- Describe the force that holds two ions together in an ionic bond.

Very strong overall attraction between positive and negative charges

- What type of energy best represents the strength of an ionic bond?

Lattice energy

- What types of bonds are present in an ionic compound that contains a polyatomic ion?

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KJ Lindholm-Leary



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Chemistry John A. Olmsted, Gregory M. Williams, Robert Charles Burk, 2016-01-14 Olmsted Burk is an introductory general chemistry text designed specifically with Canadian professors and students in mind A reorganized Table of Contents and inclusion of SI units IUPAC standards and Canadian content designed to engage and motivate readers distinguish this text from many of the current text offerings It more accurately reflects the curriculum of most Canadian institutions Instructors will find the text sufficiently rigorous while it engages and retains student interest through its accessible language and clear problem solving program without an excess of material that makes most text appear daunting and redundant

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Chemistry James N. Spencer, George M. Bodner, Lyman H. Rickard, 1999 CD ROM contains study tool illustrations photos definitions text interactive notebook drawing tools

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McMurry, Donna Jean Fredeen, Robert Pribush, 1998-01-30 Teaching "the Mole" Aina Tullberg, 1997 *Basic Principles of Inorganic Chemistry* Brian Murphy, Clair Murphy, Brian J. Hathaway, 1998 General chemistry textbooks are usually lengthy and present chemistry to the student as an unconnected list of facts In inorganic chemistry emphasis should be placed on the connections between valence shell electron configuration and the physical and chemical properties of the element *Basic Principles of Inorganic Chemistry Making the Connections* is a short concise book that emphasises these connections in particular the chemistry of the Main Group compounds With reference to chemical properties Lewis Structures stoichiometry and spider diagrams students will be able to predict or calculate the chemistry of simple polyatomic compounds from the valence shell configuration and will no longer be required to memorise vast amounts of factual chemistry This book is ideal for students taking chemistry as a subsidiary subject as well as honours degree students

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Advances in Teaching Physical Chemistry Mark David Ellison, 2008 This book brings together the latest perspectives and ideas on teaching modern physical chemistry It includes perspectives from experienced and well known physical chemists a thorough review of the education literature pertaining to physical chemistry a thorough review of advances in undergraduate laboratory experiments from the past decade in depth descriptions of using computers to aid student learning and innovative ideas for teaching the fundamentals of physical chemistry This book will provide valuable insight and information to all teachers of physical chemistry *Merrill Chemistry* Robert C. Smoot, Smoot, Richard G. Smith, Jack Price, 1998 *Chemistry and Industry*, 1996 *General Chemistry* Raymond Chang, 1986 Educational Review, 1916 **The Chemical World**, 1913 *GED Science* Cambridge, 1993-11

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