

Chapter 6 Study Guide

The Periodic Table and Periodic Law

1. Explain how Mendeleev's periodic table was in error. **Arranging the elements by mass resulted in several elements being placed in groups of elements with different properties**
How was the error fixed? **They were placed in order by increasing atomic number**
2. How was Mendeleev's periodic table organized?
In order of increasing atomic mass w/ columns with similar properties
3. What is the periodic law? **Periodic repetition of chemical and physical properties of the elements when they are arranged by increasing atomic number**
4. Identify each of the following as a metal, nonmetal, or metalloid:
 - a. Oxygen **NM**
 - b. Barium **M**
 - c. Germanium **M-oid**
 - d. Iron **M**
 - e. Neon **NM**
 - f. Tungsten **M**
5. Describe the general characteristics of metals. **Good conductors of heat & electricity, malleable (pounded into thin sheets) , & ductile (drawn into wires)**
6. Match each numbered item on the right with the lettered item that is on the left.
 1. Alkali metals **b. Group 1**
 2. Halogens **d. Group 17**
 3. Alkaline earth metals **c. Group 2**
 4. Noble gases **a. Group 18**
 - a. Group 18
 - b. Group 1
 - c. Group 2
 - d. Group 17
7. A shiny solid element also is ductile. What side of the periodic table is it likely to be found? **left**
8. What are the general properties of a metalloid? **Those of both metals and nonmetals**
List 3 metalloids. **Silicon, antimony & germanium**
9. What is the purpose of the heavy stair-step line on the periodic table? **Divides metals and nonmetals**
10. Give the chemical symbol of each of the following elements:
 - a. The two elements that are liquids are room temperature **mercury and bromine**
 - b. The noble gas with the greatest atomic mass **radon**
 - c. Any metal from group 14 **tin or lead**
 - d. Any inner transition metal **any lanthanide or actinide**
11. Explain why the radius of an atom cannot be measured directly. **Electron cloud based on probability and does not have a clearly defined edge**
12. Given any two elements within a group, is the element with the larger atomic number likely to have a larger or smaller atomic radius than the other element? **larger**
13. Which elements are characterized as having their d orbitals filled with electrons as you move left-to-right across a period? **Transition metals**
14. An element forms a negative ion when ionized. On what side of the periodic table is the element located? **right**

Chapter 6 Study Guide The Periodic Table And Law Answers

Arshad Iqbal



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