

# Chemistry 1C

## Worksheet #3

1. At 2200°C,  $K_p = 0.050$  for the reaction:



a. What is the partial pressure of NO in equilibrium with  $\text{N}_2$  and  $\text{O}_2$  that was placed in a flask at initial pressures of 0.80 and 0.20 atm?

	$\text{N}_2$	$\text{O}_2$	$2\text{NO}$
Initial	0.8	0.20	0
Change	-x	-x	+2x
Equilibrium	$0.8-x$	$0.20-x$	2x

$$K_p = (2x)^2 / ((0.80-x)(0.20-x))$$

Solve by quadratic formula and get  $x = 0.009$  atm or  $x = 0.052$ . So bc you cannot have negative pressure, the answer is 0.009

b. What happens to the concentration of the NO if we increase the temperature?

Increases

c. What happens to the concentration of  $\text{O}_2$  when you add  $\text{N}_2$  at constant volume? Decreases

d. What happens to the concentration of  $\text{O}_2$  when you add Ar at constant volume? No change

2. A 1.00 L flask was filled with 2.0 mol  $\text{SO}_2 (\text{g})$  and 2.0 mol  $\text{NO}_2 (\text{g})$  and heated.

After equilibrium was reached, it was found that 1.3 moles  $\text{NO}_2$  was present.

Assume that the reaction occurs under these conditions



a. Calculate the equilibrium constant,  $K_c$ , for the above reaction.

	$\text{SO}_2$	$\text{NO}_2$	$\text{SO}_3$	$\text{NO}$
Initial	2.0	2.0	0	0
Change	-x	-x	+x	+x
Equilibrium	$2.0-x$	$2.0-x$	+x	+x

$$K_c = \frac{(\text{mol}(\text{SO}_3))(\text{mol}(\text{NO}))}{(\text{mol}(\text{SO}_2))(\text{mol}(\text{NO}_2))} = \frac{x^2}{(2.0-x)^2}$$

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