Name:

Period:

# Atomic Structure

6 Carbon 12.01

Atomic Number

Element Symbol

Element Name

Atomic Mass

The atomic number represents these two parts of the atom:

number of protons a number of electrons

You can calculate the number of neutrons by subtracting

Atomic Mass minus Atomic Number

Nitrogen 14.01

Protons:

Electrons:

Neutrons:

Hydrogen, 1.00

Protons:

Electrons:

Neutrons:

Oxygen

16.00

Protons:

Electrons:

Neutrons:

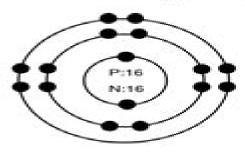
8

#### BOHR DIAGRAMS

- Draw a circle in the center representing the nucleus.
- Label the number of protons and neutrons.
- Draw the electrons in their energy levels.
  - The first energy level can hold <u>2</u> electrons
  - The second energy level can hold electrons
  - The third energy level can hold 8 electrons

The electrons in the outermost energy level are called:

Valence Electrons



Element: \_\_\_ Sulfur

Number of Valence Electrons: 6

#### LEWIS DIAGRAMS

A lewis diagram shows only the symbol with the Valence electrons





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