

Chemistry Worksheet Class 11 on Chapter 1 Some Basic Concepts of Chemistry - Set 2

Q1. What is the mass of one atom of C – 12 in grams?

- a.) 1.992×10^{-23} g
- b.) 1.989×10^{-23} g
- c.) 1.892×10^{-23} g
- d.) 1.965×10^{-23} g

Q2. Which of the following laws states that the total mass of the reactants is equal to the total mass of the products in a chemical reaction?

- a.) Law of multiple proportions
- b.) Law of conservation of mass
- c.) Avogadro's law
- d.) Law of constant proportions

Q3. The law of multiple proportions was stated by:

- a.) Dalton
- b.) Proust
- c.) Avogadro
- d.) None of the above

Q4. The units $\text{kg m}^{-1} \text{s}^{-2}$ are of ____.

- a.) Force
- b.) Pressure
- c.) Acceleration
- d.) Velocity

Q5. The empirical formula of glucose is ____.

- a.) CH_2O
- b.) CHO_2
- c.) $\text{C}_2\text{H}_2\text{O}$
- d.) CH_2O_2

Q6. Express the following in the scientific notation:

- i.) 0.0048

Chapter 2nuclear Chemistry Answers

Robert S. Boikess



Chapter 2 nuclear Chemistry Answers:

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