

# The s-, p-, d-, and f-Block Elements

Use with Chapter 6,  
Section 6.2

1. What are the four sections, or blocks, of the periodic table? s p d f
2. What does each block represent?  
an energy sublevels
3. What do elements in the s-block have in common?  
they all have 2 positions
4. What is the valence electron configuration of each element in group 1A? 1
5. What is the valence electron configuration of each element in group 2A? 2
6. Why does the s-block span two groups of elements?  
because each has 2 <sup>electrons</sup> positions
7. Why does the p-block span six groups of elements?  
because each has 6 <sup>electrons</sup> positions
8. Why are there no p-block elements in period 1?  
because helium is there
9. What is the ending of the electron configuration of each element in group 4A? ?
10. What is the electron configuration of neon? \_\_\_\_\_
11. In what period does the first d-energy sublevel appear? \_\_\_\_\_
12. Why does the d-block span ten groups of elements?  
\_\_\_\_\_
13. What is the ending of the electron configuration of each element in group 3B? \_\_\_\_\_
14. What is the electron configuration of titanium? \_\_\_\_\_
15. In what period does the first f-energy sublevel appear? \_\_\_\_\_
16. Determine the group, period, and block for the element having the electron configuration  $[\text{Xe}]4f^{14}5d^{10}6s^26p^3$ .  
a. group \_\_\_\_\_ b. period \_\_\_\_\_ c. block \_\_\_\_\_

# Chapter 6 Teaching Transparency The Periodic Table Answers

**JE Gale**



## Chapter 6 Teaching Transparency The Periodic Table Answers:

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