

1. A 33.7-g sample of HF is dissolved in water to give  $2.0 \times 10^2$  mL of solution. The concentration of the solution is:

A) 1.68 M  
B) 0.34 M  
C) 0.17 M  
**D) 8.4 M**  
E) 16.9 M

$$[HF] = \frac{33.7g HF}{2.0 \times 10^2 mL} \times \frac{1000 mL}{1 L} \times \frac{1 mol HF}{20.01g HF} = 8.4 M$$

2. What mass of calcium chloride,  $CaCl_2$ , is needed to prepare 3.419 L of a 1.56 M solution?

A) 243 g  
B) 5.33 g  
C) 50.6 g  
D) 111 g  
**E) 592 g**

$$\frac{3.419 L}{1} \times \frac{1.56 mol CaCl_2}{L} \times \frac{110.98g CaCl_2}{1 mol CaCl_2} = 592g CaCl_2$$

3. A 51.24-g sample of  $Ba(OH)_2$  is dissolved in enough water to make 1.205 liters of solution. How many mL of this solution must be diluted with water in order to make 1.000 liter of 0.1000 molar  $Ba(OH)_2$ ?

**A) 403 mL**  
B) 248 mL  
C) 24.8 mL  
D) 83.0 mL  
E) 331 mL

$$[Ba(OH)_2] = \frac{51.24g Ba(OH)_2}{1.205 L} \times \frac{1 mol Ba(OH)_2}{171.36g Ba(OH)_2} = 0.2481 M$$

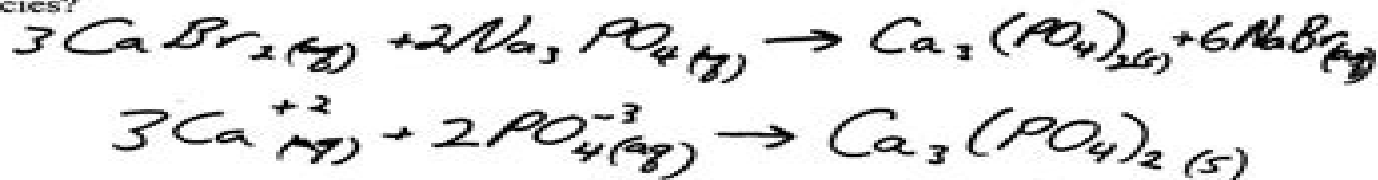
$$M_1 = 0.1000 M \quad M_2 = 0.2481 M$$

$$V_1 = 1.000 L \quad V_2 = ?$$

$$V_2 = \frac{0.1000 M \times 1.000 L}{0.2481 M} = 0.4031 L = 403 mL$$

4. The net ionic equation for the reaction of calcium bromide and sodium phosphate contains which of the following species?

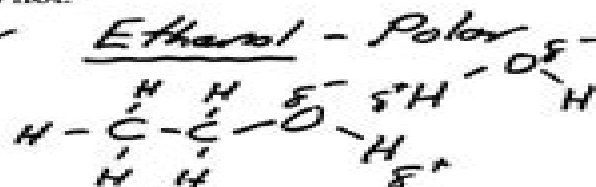
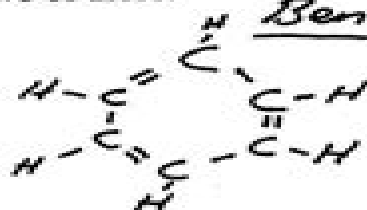
A)  $Ca^{2+}(aq)$   
B)  $PO_4^{3-}(aq)$   
C)  $2Ca_3(PO_4)_2(s)$   
D)  $6NaBr(aq)$   
**E)  $3Ca^{2+}(aq)$**



5. Consider two organic molecules, ethanol and benzene. One dissolves in water and the other does not. Why?

A) They have different molar masses.  
B) One is ionic, the other is not.  
C) One is an electrolyte, the other is not.  
**D) Ethanol contains a polar O—H bond, and benzene does not.**  
E) Two of these.

\* Know the general solubility rule of "Likes Dissolve Likes"



# Chapter 4 Chemistry Review Answers

**John Moore**



## Chapter 4 Chemistry Review Answers:

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