

## Chapter 6 Practice Test: Chemical Bonding

In the space provided, write the letter of the term or phrase that best completes each statement or best answers each question.

- \_\_\_\_\_ 1. The charge on an ion is  
a. always positive.  
b. always negative.  
c. either positive or negative.  
d. zero.
- \_\_\_\_\_ 2. According to the octet rule, a calcium atom has a tendency to  
a. lose one electron.  
b. lose two electrons.  
c. gain one electron.  
d. gain two electrons.
- \_\_\_\_\_ 3. If a compound forms by ionic bonding, which is *not* true?  
a. A positively charged atom or group of atoms attracts a negatively charged atom or group of atoms.  
b. The net charge of the compound is zero.  
c. The compound contains just two atoms, each of opposite charge.  
d. Several ions group together in a tightly packed structure.
- \_\_\_\_\_ 4. The only property listed that is *not* characteristic of ionic compounds is  
a. high melting point.  
b. hardness.  
c. lack of crystal structure.  
d. good conductor of electricity.
- \_\_\_\_\_ 5. Which formula listed below represents a polyatomic ion?  
a.  $\text{NH}_4^+$   
b.  $\text{H}_2\text{SO}_4$   
c.  $\text{NaCl}$   
d.  $\text{H}_2\text{O}$
- \_\_\_\_\_ 6. The melting points of ionic compounds are higher than the melting points of molecular compounds because  
a. ionic substances tend to vaporize at room temperature.  
b. ionic substances are brittle.  
c. attractive forces between ions are greater than the attractive forces between molecules.  
d. the numbers of positive and negative charges are equal in an ionic compound.

# Chapter 4 Test Chemical Bonds Answer Key

**Jianjun Gao**



## **Chapter 4 Test Chemical Bonds Answer Key:**

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**Chemical Interactions**, 2005 **Prentice Hall**

**Science Explorer** Michael J. Padilla, 2002 Chemistry, Student Study Guide John A. Olmsted, Gregory M.

Williams, 2002-01-04 This third edition continues to innovate by providing students with an integrated and modern approach to the subject The text emphasizes the modern tools of chemistry while incorporating historical evidence and its unique molecular quantitative emphasis is further reinforced by an integrated media package developed by the authors Also of benefit is the just in time presentation of key content only providing details once they are needed While key topics and analytical techniques have been updated there is now an additional third chapter on chemical equilibrium The authors have also developed an expanded and more integrated problem solving emphasis that now incorporates a 4 step strategy

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text chapter will then be covered with the goal of reinforcing the ideas mentioned in introductory chapters and to tie them to appropriate topics within a chapter

## Chapter 4 Test Chemical Bonds Answer Key Book Review: Unveiling the Magic of Language

In an electronic digital era where connections and knowledge reign supreme, the enchanting power of language has become more apparent than ever. Its ability to stir emotions, provoke thought, and instigate transformation is actually remarkable. This extraordinary book, aptly titled "**Chapter 4 Test Chemical Bonds Answer Key**," published by a very acclaimed author, immerses readers in a captivating exploration of the significance of language and its profound effect on our existence. Throughout this critique, we will delve into the book's central themes, evaluate its unique writing style, and assess its overall influence on its readership.

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