

Chapter 6

Chapter 6: Thermochemistry

6.2 - Enthalpy and Calorimetry

Enthalpy

$$H = E + PV$$

For a chemical reaction, enthalpy change is given by the equation

$$\Delta H = H_{\text{products}} - H_{\text{reactants}}$$

Products of a reaction have a greater enthalpy (ΔH):

- ΔH is positive
- heat is absorbed by the system (endothermic reaction)

Reactants of a reaction have a greater enthalpy (ΔH):

- ΔH is negative
- heat is released by the system (exothermic reaction)

Calorimetry

Constant-pressure calorimetry

change in enthalpy = change in heat

$$\Delta H = q_p$$

Chapter 6 Study Guide Chemistry

Dixie Goss



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traffic with a consideration of the mechanics of motor vehicles would thus be inappropriate In writing this volume I have wrestled with the question of the appropriate level at which to address the physics underlying many of the techniques used in protein isolation I have tried to strike a level as would be used by a mechanic with perhaps a slight leaning towards an engineer i e a practical level offering appropriate insight but with minimal mathematics Some people involved in biochemical research have a minimal grounding in chemistry and physics and so I have tried to keep it as simple as possible Study

Guide for Chemistry, Third Edition [by] Steven S. Zumdahl Paul B. Kelter,1993 Student Study Guide to Accompany Petrucci's General Chemistry Robert K. Wismer,1985 Introduction to Chemistry, Study Guide T. R. Dickson,1999-07-07

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