

Chem 115 POGIL Worksheet - Week 12

Molecular Shapes

Why?

Contrary to the impression that Lewis structures may give, many molecules have three-dimensional geometries. These molecular shapes are very important to understanding how molecules interact with each other, both chemically and physically. Although the Lewis structures themselves do not convey shape information, they can be used as the starting point for applying a conceptually simple but powerful approach to predicting molecular geometries. This method is called the Valence Shell Electron Repulsion Theory, or VSEPR for short. Knowing the shape of a molecule enables us to predict whether or not it has an electrical polarity, which is an important property determining how the species interacts with other molecules.

Learning Objective

- Understand the basis of the VSEPR theory.
- Know the expected geometries for one through six electron domains about a central atom
- Understand how bond pairs and lone pairs about a central atom interact to produce the molecular shape
- Know how bond polarity and molecular shape combine to make a molecule polar or nonpolar.

Success Criteria

- Be able to use a Lewis structure and VSEPR theory to predict the shape of molecules
- Be able to sketch the shapes of simple molecules
- Be able to name electron domain and molecular shape geometries of simple molecules
- Be able to predict whether a particular molecule is polar or nonpolar

Prerequisite

- Have read sections 9.1-9.3.

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