

CHEM 116-Dr. Babb's Sections
Answer Key to Exam III Lecture Problem Sheet

146. In general, any factor that weakens the bond to hydrogen, makes the substance more acidic. For binary acids within the same group such as HCl, HF, HI, and HBr, the atom bonded to hydrogen (F, Cl, Br, I) becomes larger as proceed down the periodic table and this tends to weaken the bond to H. Thus, binary acids become more acidic as proceed from top to bottom with a column. Most acidic: $\text{HI} > \text{HBr} > \text{HCl} > \text{HF}$:Least acidic; Most acidic: $\text{H}_2\text{Te} > \text{H}_2\text{S} > \text{H}_2\text{O}$:Least acidic.
147. For binary acids within the same row such as CH_4 , NH_3 , H_2O , and HF, the atom the hydrogen is bonded to (C, N, O and F) becomes more electronegative as proceed across the row and this tends to weaken the bond to H. Thus these binary acids become more acidic as proceed from left to right within a row. Most acidic: $\text{HF} > \text{H}_2\text{O} > \text{NH}_3 > \text{CH}_4$:Least acidic.
148. In oxyacids, the H is always bonded to one of the oxygens. Any factor that weakens the O-H bond, makes the substance more acidic. Within an oxyacid series such as HClO , HClO_2 , HClO_3 , and HClO_4 , as the number of oxygens bonded to the central atom increases, the oxidation number of the central atom increases causing a weakening of the O-H bond strength and an increase in the acidity. Most acidic: $\text{HClO}_4 > \text{HClO}_3 > \text{HClO}_2 > \text{HClO}$.

For a series of oxyacids with the same number of oxygens but with different central atoms such as HOBr , HOCl , HOI , as the electronegativity of the central atom increases, the O-H bond strength weakens and the acidity increases. Most acidic: $\text{HOCl} > \text{HOBr} > \text{HOI}$:Least acidic.

149. Compare strengths of acids on reactant and product sides. The stronger acid has a greater tendency to give up its proton; therefore, the equilibrium will lie on the opposite side from the strongest acid.
- A. HBr is a stronger acid than HF. Eq lies on side opposite the HBr (i.e. on reactant side); thus, $[\text{Reactants}] > [\text{Products}]$ and $K < 1$.
 - B. H_2O is a stronger acid than NH_3 . Eq lies on side opposite the H_2O (i.e. on product side); thus, $[\text{Products}] > [\text{Reactants}]$ and $K > 1$.
 - C. HClO_2 is a stronger acid than HClO . Eq lies on side opposite the HClO_2 (i.e. on reactant side); thus, $[\text{Reactants}] > [\text{Products}]$ and $K < 1$.
 - D. H_3O^+ is a stronger acid than H_2O . Eq. lies on side opposite the H_3O^+ (i.e. on the reactant side); thus, $[\text{Reactants}] > [\text{Products}]$ and $K < 1$.
 - E. HBrO_3 is a stronger acid than HIO_3 . Eq. lies on side opposite the HBrO_3 (i.e. on the product side); thus, $[\text{Products}] > [\text{Reactants}]$ and $K > 1$.
154. As solid $\text{NaC}_2\text{H}_3\text{O}_2$ is added to 0.500 M $\text{HC}_2\text{H}_3\text{O}_2$, the $[\text{H}_3\text{O}^+]$ decreases and the pH increases. Why? Look at effect of addition of $\text{NaC}_2\text{H}_3\text{O}_2$, a strong electrolyte, on the acetic acid eq.
- $$\text{HC}_2\text{H}_3\text{O}_2(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{C}_2\text{H}_3\text{O}_2^-(\text{aq}) + \text{H}_3\text{O}^+(\text{aq})$$
- $$\text{NaC}_2\text{H}_3\text{O}_2(\text{s}) \rightarrow \text{Na}^+(\text{aq}) + \text{C}_2\text{H}_3\text{O}_2^-(\text{aq})$$
- As $\text{NaC}_2\text{H}_3\text{O}_2$ is added to an eq. mixture of 0.500 M $\text{HC}_2\text{H}_3\text{O}_2$, the common ion $\text{C}_2\text{H}_3\text{O}_2^-$ is added which shifts the acetic acid eq. to the left decreasing the $[\text{H}_3\text{O}^+]$ and increasing the pH.
- Addition of NaCl to an eq. mixture of 0.500 M $\text{HC}_2\text{H}_3\text{O}_2$ has no effect on the pH or acetic acid eq. because NaCl has no ion in common with the acetic acid eq.

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