

THE "OFFICIAL" CHEMISTRY 12 REDOX & ELECTROCHEMISTRY STUDY GUIDE

- DO ALL THE QUESTIONS** in this booklet. These are actual Provincial Exam questions! Your own provincial exam and unit test will include questions similar to the ones in this booklet!
- RESIST THE URGE TO LOOK AT THE ANSWER KEY** until you have given all the questions in the section your best effort. Don't do one question, then look at the key, then do another and look at the key, and so on. Each time you look at one answer in the study guide, your eye will notice other answers around them, and this will reduce the effectiveness of those questions in helping you to learn.
- LEARN FROM YOUR MISTAKES!** If you get a question wrong, **figure out why!** If you are having difficulty, **talk to your study partner**, or maybe **phone someone in your Peer Tutoring group**. Get together with group members or other students from class and work on these questions together. Explain how you got your answers to tough questions to others. In explaining yourself to someone else, you will learn the material better yourself (try it)! Ask your teacher to explain the questions to you during tutorial or after school. **Your goal should be to get 100% on any Chemistry 12 multiple choice test** - learning from your mistakes in this booklet will really help you in your efforts to meet this goal!
- This is REALLY CRUCIAL: DO NOT mark the answer anywhere on the questions themselves.** For example, do not circle any of options A B C or D - instead use a different sheet of paper to place your answers on. By avoiding this urge, you can re-use this study guide effectively again, when preparing for your final exam. In the box to the left, put an asterisk or small note to yourself to indicate that you got the question wrong and need to come back to it. If you got the question correct initially, a check mark might be assurance that you understand this type of question and therefore can concentrate on other questions that present a challenge to you.
- Check Off the STATUS box on the PRESCRIBED LEARNING OUTCOMES sheet.** I have tried to organize the questions in the identical sequence to which they appear on your Redox/Electrochemistry Prescribed Learning Outcome sheet. By doing this, you can be confident that you know everything you need to know for both the UNIT EXAM and PROVINCIAL EXAM!

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INTRODUCTION

1. [501](#) Which one of the following is the strongest reducing agent?

A. Γ^- B. F^- C. Cl^- D. Br^-

2. [501](#) In the reaction below:



the oxidizing agent is

A. I_2 B. Γ^- C. H^+ D. ClO_3^-

3. [501](#) Which of the following is the strongest reducing agent?

- A. cobalt
B. copper
C. calcium
D. chromium

4. [501](#) An oxidizing agent is

- A. reduced as it loses electrons.
B. reduced as it gains electrons.
C. oxidized as it loses electrons.
D. oxidized as it gains electrons.

5. [501](#) Which of the following is the strongest oxidizing agent?

- A. Mn
B. Mn^{2+}
C. acidified MnO_2
D. acidified MnO_4^-

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Bioactive Natural Products (Part C) Atta-ur Rahman, 2000-10-20 Natural products play an integral and ongoing role in promoting numerous aspects of scientific advancement and many aspects of basic research programs are intimately related to natural products. The significance therefore of the Studies in Natural Product Chemistry series edited by Professor Atta-ur Rahman cannot be overestimated. This volume in accordance with previous volumes presents us with cutting edge contributions of great importance. Analytical Pyrolysis Handbook Karen D. Sam, Thomas P. Wampler, 2021-03-22

Analytical pyrolysis allows scientists to use routine laboratory instrumentation for analyzing complex opaque or insoluble samples more effectively than other analytical techniques alone. Analytical Pyrolysis Handbook Third Edition is a practical guide to the application of pyrolysis techniques to various samples and sample types for a diversity of fields including microbiology, forensic science, industrial research, and environmental analysis. The much anticipated third edition incorporates recent technological advances that increase the technique's sensitivity to trace elements, improve its reproducibility, and expand its applicability. The book reviews the types of instrumentation available to perform pyrolysis and offers guidance for interfacing instruments and integrating other analytical techniques including gas chromatography and mass spectrometry. Fully updated with new sample pyrograms, figures, references, and real world examples, this edition also highlights new areas of application including cultural materials, forensic analysis, and environmental studies. This book illustrates how the latest advances make pyrolysis a practical, cost-effective, reliable, and flexible alternative for increasingly complex sample analyses. Analytical Pyrolysis Handbook Third Edition is an essential one-stop guide for determining if pyrolysis meets application specific needs as well as performing pyrolysis and handling the data obtained. Applied Pyrolysis Handbook Thomas P. Wampler, 2006-12-13 Analytical pyrolysis allows scientists to use routine laboratory instrumentation for analyzing complex opaque or insoluble samples more effectively than other analytical techniques alone. Applied Pyrolysis Handbook Second Edition is a practical guide to the application of pyrolysis techniques to various samples and sample types for a dive

Cumulated Index Medicus, 1994 **Agricultural Reviews**, 1990 Lunar and Planetary Science, 1992 **Indian Journal of Applied Chemistry**, 1969 Chemical Analysis Michael Raymond Frederick Ashworth, 1964 Canadian Journal of Chemistry, 2007 **Journal of the Indian Chemical Society** Indian Chemical Society, Calcutta, 1981-07

Journal of the Mysore University, 1967 **Compounds with Fluorine and Chlorine** Reinhard Haubold, Peter Merlet, Jörn v. Jouanne, Hannelore Keller-Rudek, Ulrike Ohms-Bredemann, Carol Strametz, Joachim Wagner, Astrid Wietelmann, 1991-11-26 The volume describes all compounds that consist of bromine and fluorine and/or chlorine and may additionally contain noble gases, hydrogen, oxygen, and nitrogen. The description of chemical and physical properties of binary compounds between bromine and fluorine takes up most of the volume because this class of compounds includes BrF₃ and BrF₅ which have considerable technical interest. Especially the oxidizing and fluorinating properties of BrF₃ make it a

convenient reactant for the preparation of inorganic fluorides On the other hand the diatomic molecule BrF is well characterized by spectroscopic methods but its chemistry is less known because of its instability Other neutral species such as Br₂F Br₂F₂ BrF₂ and BrF₆ only exist in matrices at low temperatures and the existence of BrF₄ and BrF₇ is even doubted Some of the ions including BrF₂⁺ BrF₂⁻ Br₃F₁₀⁺ BrF₄⁺ BrF₄⁻ BrF₆⁺ and BrF₆⁻ can be stabilized as salts **Journal of the Indian Chemical Society** Indian Chemical Society,1981 **From Mantle to Meteorites** K. Gopalan,1990 Professor Lai combines a deep understanding of the physics of high energy particle interactions with a rare insight into the evolution of the tenuous trails which they leave behind in naturally occurring chemical systems Thus by tracking down the exotic products of these high energy interactions into the specific niches of various low energy chemical reservoirs he has made pioneering contributions to such diverse fields a astrophysics planetary evolution oceanography and archaeology Journal of Applied Chemistry ,1962 **Quarterly Journal of the Chemical Society of London** Chemical Society (Great Britain),1971 **Energy Research Abstracts** ,1987 **Russian Chemical Reviews** ,1989 **Index Medicus** ,2001 Vols for 1963 include as pt 2 of the Jan issue Medical subject headings **Newer Redox Titrants** Antonín Berka,Jaroslav Vulterin,Jaroslav Zýka,1965

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