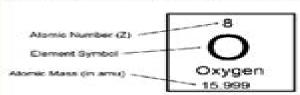
#### ACS General Chemistry Cheat Sheet #1

Protons: positively charged atoms Neutrons: neutral charged atoms Electrons: negatively charged atoms

Is otopie: an atom with the same number of protons, but different number of neutrons.



NOTE: you need to remember the atomic number and atomic mass of an element to calculate certain things.

#### Questions from the ACS General Chemistry-Second Edition Exam Book

An atom of strontium-90, 10 Sr contains,

a. 38 electrons, 38 protons, and 52 neutrons b. 38 electrons, 38 protons, 90 neutrons c. 52 electrons, 52 protons, 38 neutrons d. 52 electrons, 38 protons, 38 neutrons

The correct answer for this question would be a. The atomic mass number is the sum of protons and neutrons. The atomic number is the amount of neutrons a specific element has. In this case, we have 38 protons which is the same amount for electrons, so we can go ahead and eliminate choices c and d. If you take the atomic mass number and subtract it from the atomic number, you can get the number of neutrons the element has.

#### "One thing you do need to know for the exam is how to calculate relative abundance."

Chlorine occurs naturally as a mixture of two isotopes with atomic masses of 34.97 amu and 36.97 amu. What are the relative abundances of these isotopes?

Remember: Isotopies are elements with the same number of protons, but different number of neutrons.

The formula that can be used to calculate the answer is  $[(mass\ isotope_1)^{\frac{96\ isotope_1}{100\%}}] + [(mass\ isotope_2)^{\frac{96\ isotope_2}{100\%}}] = average\ atomic\ mass$ 35.45 amu - 34.97 amu

100%-24%=76%

The correct answer is 76%.

# **Chem 1212 Acs Final Exam Study Guide**

**LUCY T. EUBANKS** 

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