

Chem 1411 Study Guide for Exam 2

I. Things that you need to know !!!!

1. Evidence of chemical reaction: precipitation, color change, gas evolution, neutralization, heat, etc.
2. Definition of electrolytes: strong, weak, and non-electrolytes.
3. Chemical equation: molecular equation, ionic equation, net ionic equation, spectator ions.
4. Precipitation reaction
5. Acid-base neutralization reaction: Definition of acids & bases and their strong and weak ones.
6. Oxidation-reduction reactions: definition of oxidation number and how to assign it on each atom.
7. Dilution: $M_1V_1=M_2V_2$
7. Properties of light: wavelength, frequency, and speed of light.
8. Quantum numbers: i) principal quantum number (shell number); ii) angular momentum quantum number (subshell number); iii) magnetic quantum number; iv) spin quantum number
9. Shape of orbital
10. Electron configurations and their corresponding rules: Aufbau, Pauli, and Hund rules.
11. Electron diagram
12. The first law of Thermodynamics
13. State function
14. Heat(q) or Enthalpy (ΔH) calculation (Hess Law)
15. Percent yield, limiting agent (Chap 3)

II. Sample questions

1. How many moles of NO_2 are needed to form 5.24 moles of NO if there is plenty of water present?



- A) 15.72 moles NO_2 B) 25.3 moles NO_2 C) 8.44 moles NO_2
D) 5.63 moles NO_2 E) 1.83 moles NO_2

2. How many grams of NO are formed when 6.21 moles of NO_2 are reacted?



- A) 19.9 g NO B) 27.9 g NO C) 62.1 g NO D) 3.59 g NO E) 11.7 g NO

3. If 14.01 g sample of N_2 reacts with 4.02 g of H_2 to form ammonia (NH_3), what is the mass of ammonia formed?

- A) 34.02 g B) 11.10 g C) 17.01 g D) 30.02 g E) 23.07 g

4. The chemical equation shows that carbonic acid is decomposed to form water and carbon dioxide by heating. What is the mass of carbon dioxide formed from 12.40 g of carbonic acid?



- A) 7.58 g B) 8.80 g C) 2.20 g D) 6.20 g

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Chemistry 1411 Enrique Olivas, 2014-08-19 THE MAIN OBJECTIVES DURING THE DEVELOPMENT OF THIS BOOK WAS TO BETTER PREPARE STUDENTS THAT NEED TO TAKE THE FINAL COMPREHENSIVE AMERICAN CHEMICAL SOCIETY ACS EXAM AS WELL AS THOSE STUDENTS THAT SEEK ADMISSION IN MEDICAL AND PHARMACY SCHOOLS THE STUDY GUIDE DESCRIBES IN AN OUTLINE FORMAT THE MOST IMPORTANT TOPICS COVERED IN GENERAL CHEMISTRY I THE BOOK SUMMARIZES THE MAIN OBJECTIVES THAT STUDENTS ENROLLED IN THE COURSE SHOULD LEARN THE BOOK GIVES MANY SAMPLE PROBLEMS WITH STEP WISE SOLUTIONS SO THAT STUDENTS CAN FOLLOW THE MATERIAL EASIER WHEN NECESSARY MATHEMATICAL FORMULAS ARE GIVEN ALL THROUGHOUT TO FACILITATE THE SOLUTIONS TO NUMERICAL PROBLEMS THIS STUDY GUIDE CAN BE USED BY ANY COLLEGE STUDENT ENROLLED IN GEN CHEM I REGARDLESS OF THE TEXT USED ONE OF THE PITFALLS OF MOST TEXTS IS THE EXCESSIVE AMOUNT OF MATERIAL COVERED BUT FAIL TO EMPHASIZE THE MOST IMPORTANT FEATURES OF THE TOPIC COVERED

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Guang Yang, Muhammad Wajid Ullah, Orlando Rojas, Ronan McCarthy, 2021-11-18 Drs Ullah and Yang hold patents related to cellulose material All other Topic Editors declare no competing interests with regard to the Research Topic subject This Research Topic is dedicated to Prof Lina Zhang on the occasion of her 80th Birthday in gratitude esteem and affection

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Petroleum Abstracts, 1993

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Table of Contents **Chem 1411 Acs Study Guide**

1. Understanding the eBook Chem 1411 Acs Study Guide
 - The Rise of Digital Reading Chem 1411 Acs Study Guide
 - Advantages of eBooks Over Traditional Books
2. Identifying Chem 1411 Acs Study Guide
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chem 1411 Acs Study Guide
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chem 1411 Acs Study Guide
 - Personalized Recommendations
 - Chem 1411 Acs Study Guide User Reviews and Ratings
 - Chem 1411 Acs Study Guide and Bestseller Lists

5. Accessing Chem 1411 Acs Study Guide Free and Paid eBooks
 - Chem 1411 Acs Study Guide Public Domain eBooks
 - Chem 1411 Acs Study Guide eBook Subscription Services
 - Chem 1411 Acs Study Guide Budget-Friendly Options
6. Navigating Chem 1411 Acs Study Guide eBook Formats
 - ePub, PDF, MOBI, and More
 - Chem 1411 Acs Study Guide Compatibility with Devices
 - Chem 1411 Acs Study Guide Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chem 1411 Acs Study Guide
 - Highlighting and Note-Taking Chem 1411 Acs Study Guide
 - Interactive Elements Chem 1411 Acs Study Guide
8. Staying Engaged with Chem 1411 Acs Study Guide
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chem 1411 Acs Study Guide
9. Balancing eBooks and Physical Books Chem 1411 Acs Study Guide
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chem 1411 Acs Study Guide
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chem 1411 Acs Study Guide
 - Setting Reading Goals Chem 1411 Acs Study Guide
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chem 1411 Acs Study Guide
 - Fact-Checking eBook Content of Chem 1411 Acs Study Guide
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
 - Interactive and Gamified eBooks

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