

PROBLEM SET / CHEM 1411 — Ch 4: Chemical Reactions

Name \_\_\_\_\_

- 1) A 31.5 mL aliquot of  $\text{H}_2\text{SO}_4$  (aq) of unknown concentration was titrated with 0.0134 M NaOH (aq). It took 23.9 mL of the base to reach the endpoint of the titration. What was the concentration (M) of the acid?
  - A) 0.227
  - B) 0.102
  - C) 0.0051
  - D) 0.0204
  - E) 0.0102
- 2) A stock solution of  $\text{HNO}_3$  is prepared and found to contain 13.5 M of  $\text{HNO}_3$ . If 25.0 mL of the stock solution is diluted to a final volume of 0.500 L, the concentration of the diluted solution is \_\_\_\_\_ M.
  - A) 675
  - B) 0.270
  - C) 0.675
  - D) 270
  - E) 1.48
- 3) What is the concentration of iodide ions in a 0.193 M solution of barium iodide?
  - A) 0.0965 M
  - B) 0.386 M
  - C) 0.193 M
  - D) 0.0643 M
  - E) 0.579 M
- 4) A strong electrolyte is one that \_\_\_\_\_ completely in solution.
  - A) disappears
  - B) dissolves
  - C) decomposes
  - D) reacts
  - E) ionizes
- 5) There are \_\_\_\_\_ mol of bromide ions in 0.500 L of a 0.300 M solution of  $\text{AlBr}_3$ .
  - A) 0.150
  - B) 0.450
  - C) 0.167
  - D) 0.0500
  - E) 0.500
- 6) What is the concentration (M) of an aqueous methanol produced when 0.200 L of a 2.00 M solution was diluted to 0.800 L?
  - A) 8.00
  - B) 0.500
  - C) 0.800
  - D) 0.200
  - E) 0.400
- 7) Oxidation and \_\_\_\_\_ mean essentially the same thing.
  - A) activity
  - B) metathesis
  - C) corrosion
  - D) reduction
  - E) decomposition
- 8) When 0.500 mol of  $\text{HC}_2\text{H}_3\text{O}_2$  is combined with enough water to make a 300 mL solution, the concentration of  $\text{HC}_2\text{H}_3\text{O}_2$  is \_\_\_\_\_ M.
  - A) 3.33
  - B) 0.00167
  - C) 0.835
  - D) 1.67
  - E) 0.150
- 9) With which of the following will ammonium ion form an insoluble salt?
  - A) sulfate
  - B) chloride
  - C) carbonate
  - D) sulfate and carbonate
  - E) none of the above
- 10) When aqueous solutions of  $\text{AgNO}_3$  and NaI are mixed, AgI precipitates. The balanced net ionic equation is \_\_\_\_\_
  - A)  $\text{Ag}^+ (\text{aq}) + \text{I}^- (\text{aq}) \rightarrow \text{AgI} (\text{s})$
  - B)  $\text{Ag}^+ (\text{aq}) + \text{NO}_3^- (\text{aq}) \rightarrow \text{AgNO}_3 (\text{s})$
  - C)  $\text{Ag}^+ (\text{aq}) + \text{NO}_3^- (\text{aq}) \rightarrow \text{AgNO}_3 (\text{aq})$
  - D)  $\text{AgNO}_3 (\text{aq}) + \text{NaI} (\text{aq}) \rightarrow \text{AgI} (\text{aq}) + \text{NaNO}_3 (\text{s})$
  - E)  $\text{AgNO}_3 (\text{aq}) + \text{NaI} (\text{aq}) \rightarrow \text{AgI} (\text{s}) + \text{NaNO}_3 (\text{aq})$

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**Jesse Yeh**



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American Journal of Medical Technology ,1982      *A Laboratory Manual of Analytical Methods of Protein Chemistry, Including Polypeptides* Peter Alexander,Richard Joseph Block,Harold Palmer Lundgren,1960      *A Laboratory Manual of Analytical Methods of Protein Chemistry* ,1960      **Journal of the American Chemical Society** American Chemical Society,1914      **Bulletin** Omaha Public Library,1901      **Manual on Industrial Water** American Society for Testing Materials. Committee D-19 on Industrial Water,1954      **Chemical Engineering Catalog** ,1920      **Chemistry 1411** Jesse Yeh,2012-07-20      **Association Medical Journal** ,1914      Industrial Chemistry Allen Rogers,1925      **Scientific and Technical Aerospace Reports** ,1976 Lists citations with abstracts for aerospace related reports obtained from world wide sources and announces documents that have recently been entered into the NASA Scientific and Technical Information Database      **Selected Water Resources Abstracts** ,1981      Publishers' circular and booksellers' record ,1874

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