NAME ____BLUE/ GREEK/ RED __ (QUESTIONS MAY BE RANDOMIZED) SECTION ____

Calculate the change in internal energy (ΔE) for a system that is giving off 23.5/32.5/52.3 kJ of heat and is changing from 12.00/15.00/20.00 L to 6.00 L in volume at 1.50 atm pressure.
 (Given: 101.3 J = 1 L.atm)

A) -22.9 kJ	A) -33.4 kJ	A) +53.7 kJ
B) -24.1 kJ	B) -31.6 kJ	B) -53.7 kJ
C) $+24.1 \text{ kJ}$	C) +33.4 kJ	C) -50.9 kJ
D) +22.9 kJ	D) 912 kJ	D) +50.9 kJ
E) 608 kJ	E) +31.6 kJ	E) 1.42 kJ

2) A sample of copper absorbs 43.6/36.4/63.4 kJ of heat, resulting in a temperature rise of 55.5 °C, determine the mass (in kg) of the copper sample. [specific heat of copper is 0.385 J/g°C.]

A) 0.49 kg	A) 1.70 kg	A) 1.51 kg
B) 2.97 kg	B) 2.04 kg	B) 2.97 kg
C) 1.26 kg	C) 0.58 kg	C) 1.70 kg
D) 1.70 kg	D) 2.97 kg	D) 2.04 kg
E) 2.04 kg	E) 3.64 kg	E) 0.34 kg

3. Which of the following elements has the smallest atomic size?

A) Na	B) Ar	C) N	D) Ca	E) NE

- 3. Which of the following elements has the largest atomic size?

 A) Na

 B) Ar

 C) K

 D) Ca

 E) Kr
- 3. Which of the following elements has the largest atomic size?

 A) S

 B) Ca

 C) Po

 D) Ba

 E) Rn

4) What is the minimum energy <u>required</u> to decompose 465/654/564 g of PCI₃, according to the reaction below?

4 PCl3(g) → F	4(s) + 6 Cl ₂ (g)	$\Delta H^{\circ}_{TXB} =$	+1207 kJ
A) $2.31 \times 10^{3} \text{ kJ}$	A) 2.31×10^{3}	kJ	A) $2.31 \times 10^{3} \text{ kJ}$
B) 4.09 × 103 kJ	B) 4.33 × 103	kJ	B) $4.96 \times 103 \text{ kJ}$
C) $1.02 \times 10^3 \text{ kJ}$	C) 1.02×10^{3}	kJ	C) $1.02 \times 10^{3} \text{ kJ}$
D) $1.68 \times 10^{3} \text{ kJ}$	D) 1.44 × 103	Scill .	D) $1.68 \times 10^{3} \text{ kJ}$
E) $5.95 \times 10^{3} \text{ kJ}$	E) 5.75×10^3	kJ	E) $1.24 \times 10^3 \text{ kJ}$

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Yan Bai

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