

Answer Key

Question 1: A 1.12 M solution of citric acid ($\text{H}_3\text{C}_6\text{H}_5\text{O}_7$) in water has a density of 1.08 g/cm³. Calculate the mass percent, molality, and mole fraction of the citric acid.

Assume 1 L of solution, \therefore 1.12 moles Citric Acid & Mass of solution is 1080 grams

Molar Mass is $192.123 \frac{\text{g}}{\text{mol}}$

\rightarrow Mass is $MM \times N = 215.178 \text{ grams}$

Mass % = $\frac{215.178 \text{ g}}{1080 \text{ g}} = 19.9\%$

Mass of $\text{H}_2\text{O} = 1080 \text{ g} - 215.178 \text{ g} = 0.86482 \text{ kg}$

Molality = $\frac{1.12 \text{ moles}}{0.86482 \text{ kg}} = 1.3 \frac{\text{mol}}{\text{kg}}$

moles $\text{H}_2\text{O} = \frac{\text{mass}}{\text{MM}} = \frac{864.8 \text{ g}}{18.0158} = 48.005 \text{ mol}$

$X_{\text{CA}} = \frac{1.12}{48.005 + 1.12} = 0.023$

Question 2: Pure solvent X boils at 81.1°C. When 11.00 g of solute A (non-electrolyte, MM = 88.1 g/mol) is dissolved in 150.0 g of solvent X, the new boiling point is 83.5°C. When 4.00 g of solute B (non-electrolyte) is dissolved in 50.0 g of solvent X, the solution boils at 82.4°C. What is the molar mass of solute B?

A: $\Delta B_p = 2.4^\circ\text{C}$

$V_{\text{Hfr}} = 1$

Molality = $\frac{\text{moles}}{\text{kg solvent}} = \frac{(11 \text{ g} / 88.1 \text{ g/mol})}{0.150 \text{ kg}} = 0.8324 \frac{\text{mol}}{\text{kg}}$

$2.4 = \Delta B_p = 1 \cdot 0.8324 \frac{\text{mol}}{\text{kg}} \cdot K_b$

$K_b = \frac{2.4^\circ\text{C}}{0.8324 \frac{\text{kg mol}}{\text{kg}}} \rightarrow 2.8833 \frac{\text{kg}^\circ\text{C}}{\text{mol}}$

B: $\Delta B_p = 1.3^\circ\text{C} = 1 \cdot X \cdot 2.8833 \frac{\text{kg}^\circ\text{C}}{\text{mol}}$

$X = \text{molality} = \frac{1.3^\circ}{2.8833} = 0.45088 \frac{\text{mol}}{\text{kg}}$

Molality = $\frac{\text{mol}}{\text{kg solvent}} = \frac{(m / \text{MM})}{\text{kg solvent}} \therefore \text{MM} = \frac{m}{\text{mass solvent} \times \text{molality}} = \frac{4.00 \text{ g}}{0.050 \text{ kg} \times 0.45088} = 1.8 \times 10^3 \text{ g/mol}$

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