Thermochemistry Practice Problems ANSWER KEY

1. Calculate the standard enthalpy of reaction for the process

 Calculate the heat evolved in the combustion of ethyl alcohol, ΔH_{comb} C₂H₅ OH₁₈ + 3 O_{2 los} → 2 CO_{2 los} + 3 H₂O₁₈ ΔH = ?

2 C _{graphse} + 3 H _{2 lol} + ½ O _{2 lol}	→ C ₂ H ₅ OH ₂₁	$\Delta H = -277.7 kJ$
Cgraphes + O2 kg → CO2 kg		$\Delta H = -393.5 \text{ kJ}$
H _{2 (q)} + ½ O _{2 (q)} → H ₂ O (t)		ΔH = -285.8 kJ
		$\Delta H = -1366.7 \text{ kJ}$

Calculate the heat required to decompose limestone, CaCO₃, to lime, CaO, and carbon dioxide, with all substances at standard conditions.
 CaCO_{3 N} → CaO_N + CO_{2 N}

ΔH*₁ CaCO_{3 |q|} = -1206.9 kJ/mol ΔH*₁ CaO_{3 |q|} = -635.1 kJ/mol ΔH*₁ CO_{2 |q|} = -393.5 kJ/mol

 $\Delta H = 178.3 \text{ kJ}$

-46.1 kJ/mel

4. $CH_{E(g)} + 3Cl_{E(g)} \rightarrow CHCl_{D(g)} + 3HCl_{(g)} \Delta H = ?$

 Cyanamide, CH₂N₂, is a weak acid that is sometimes used as a fertilizer. Calculate the standard enthalpy of formation for cyanamide, given the following standard enthalpies of reaction:

The standard molar enthalpy change is -905.2 kJ for the oxidation of ammonia.
 4 NH_{3 kgi} + 5 O_{2 kgi} → 4 NO kgi + 6 H₂O_{3gi}
 Calculate the standard molar enthalpy of formation for ammonia based on the following standard enthalpies of formation: ΔH*₁ NO kgi = 90.3 kJ/mol and ΔH*₁ H₂O kgi = -241.8 kJ/mol

7. The thermite reaction: $2 \text{ Al}_{|\mathbb{N}|} + \text{Fe}_2 O_{3 |\mathbb{N}|} \rightarrow \text{Al}_2 O_{3 |\mathbb{N}|} + 2 \text{ Fe}_{|\mathbb{N}|} + 352 \text{ kJ}$

If you begin with 10.0 g of Al and excess Fe₂O₃, how many kJ of heat are evolved at constant pressure?

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