

Chemical bonding short notes

(1)

IONIC BOND

* 1) * polarisation \propto size of anion

$\propto \frac{1}{\text{size of cation}}$

\propto magnitude of charge

* polarisation \propto covalent character $\propto \frac{1}{\text{Ionic character}}$

\propto Ionic potential

\propto Intensity of colour

$\propto \frac{1}{\text{m.p.t. \& B.p.t.}}$

2) Solubility of ionic compounds in H_2O :

Case (i): common ions: Na^+ , Li^+ , O^{2-} , F^- , OH^- & IIA gp cations.

By \uparrow in size of ~~cat~~ ions } solubility \uparrow es.
LE \downarrow es

ex:— $Be(OH)_2 < Mg(OH)_2 < Ca(OH)_2 < Sr(OH)_2 < Ba(OH)_2$

Case (ii): common ions: Cs^+ , Rb^+ , Br^- , I^- , CO_3^{2-} & SO_4^{2-} .

By \uparrow in size of ions } solubility \downarrow es.
HE \downarrow es

ex:— $BeCO_3 > MgCO_3 > CaCO_3 > SrCO_3 > BaCO_3$

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Table of Contents Chem Fax Teachers Notes

1. Understanding the eBook Chem Fax Teachers Notes
 - The Rise of Digital Reading Chem Fax Teachers Notes
 - Advantages of eBooks Over Traditional Books
2. Identifying Chem Fax Teachers Notes
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Chem Fax Teachers Notes
 - User-Friendly Interface
4. Exploring eBook Recommendations from Chem Fax Teachers Notes
 - Personalized Recommendations
 - Chem Fax Teachers Notes User Reviews and Ratings
 - Chem Fax Teachers Notes and Bestseller Lists
5. Accessing Chem Fax Teachers Notes Free and Paid eBooks
 - Chem Fax Teachers Notes Public Domain eBooks
 - Chem Fax Teachers Notes eBook Subscription Services
 - Chem Fax Teachers Notes Budget-Friendly Options

6. Navigating Chem Fax Teachers Notes eBook Formats
 - ePub, PDF, MOBI, and More
 - Chem Fax Teachers Notes Compatibility with Devices
 - Chem Fax Teachers Notes Enhanced eBook Features
7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Chem Fax Teachers Notes
 - Highlighting and Note-Taking Chem Fax Teachers Notes
 - Interactive Elements Chem Fax Teachers Notes
8. Staying Engaged with Chem Fax Teachers Notes
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Chem Fax Teachers Notes
9. Balancing eBooks and Physical Books Chem Fax Teachers Notes
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Chem Fax Teachers Notes
10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
11. Cultivating a Reading Routine Chem Fax Teachers Notes
 - Setting Reading Goals Chem Fax Teachers Notes
 - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Chem Fax Teachers Notes
 - Fact-Checking eBook Content of Chem Fax Teachers Notes
 - Distinguishing Credible Sources
13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
14. Embracing eBook Trends
 - Integration of Multimedia Elements

- Interactive and Gamified eBooks

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