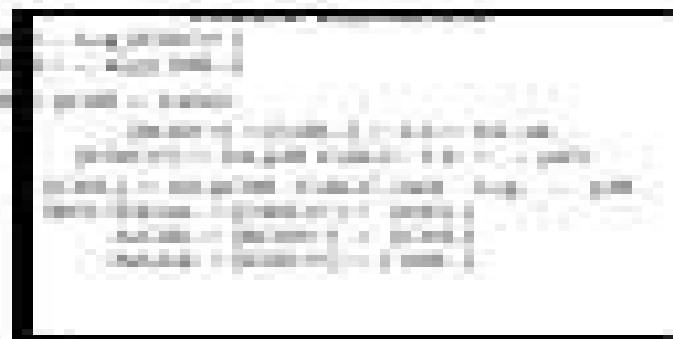
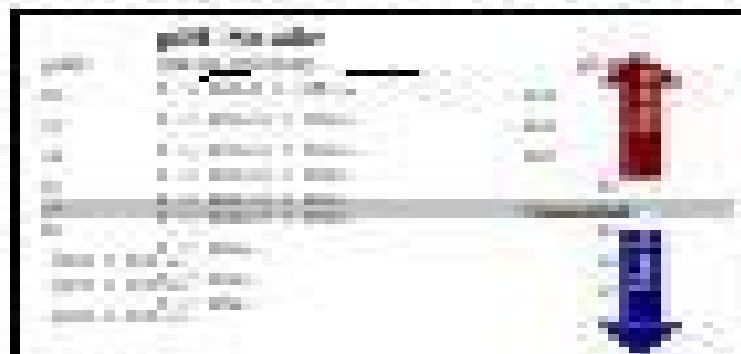


The pH Scale

Chem. 101/102, Spring 2018

Name: _____

An acidic solution has a pH less than 7.0, a neutral solution has a pH of 7.0, and a basic solution has a pH greater than 7.0. The pH scale is logarithmic, meaning that the value decreases by a factor of 10 for each unit increase in pH. For example, a solution with a pH of 3 is 10 times more acidic than a solution with a pH of 4, and 100 times more acidic than a solution with a pH of 5.



Example 1

What is the pH of a solution with $[H^+] = 1.0 \times 10^{-3} M$?

$pH = -\log [H^+]$
 $pH = -\log (1.0 \times 10^{-3})$
 $pH = 3.00$

What is the pH of a solution with $[OH^-] = 1.0 \times 10^{-4} M$?

$pOH = -\log [OH^-]$
 $pOH = -\log (1.0 \times 10^{-4})$
 $pOH = 4.00$

Since $pH + pOH = 14.00$, $pH = 14.00 - 4.00 = 10.00$

or $pH = 10.00$

What is the pH of a solution with $[H^+] = 1.0 \times 10^{-10} M$?

$pH = -\log [H^+]$
 $pH = -\log (1.0 \times 10^{-10})$
 $pH = 10.00$

What is the pH of a solution with $[OH^-] = 1.0 \times 10^{-5} M$?

$pOH = -\log [OH^-]$
 $pOH = -\log (1.0 \times 10^{-5})$
 $pOH = 5.00$

Since $pH + pOH = 14.00$, $pH = 14.00 - 5.00 = 9.00$

Rank the following solutions. Which will be the most acidic?

1. A solution with a pH of 3.0
2. A solution with a pH of 4.0
3. A solution with a pH of 5.0
4. A solution with a pH of 6.0
5. A solution with a pH of 7.0
6. A solution with a pH of 8.0
7. A solution with a pH of 9.0
8. A solution with a pH of 10.0
9. A solution with a pH of 11.0
10. A solution with a pH of 12.0
11. A solution with a pH of 13.0
12. A solution with a pH of 14.0

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Chem Worksheet 19 3 Answers

**A. Kayode Coker, Rahmat Sotudeh-
Gharebagh**

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- Chapter 3: Chem Worksheet 19 3 Answers in Everyday Life
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