

Worksheet 5

- 1) How many moles of Cl^- are in 518 mL of 0.27 M FeCl_3 ? $\frac{3 \text{ mol Cl}^-}{1 \text{ mol FeCl}_3} \times 0.14 \text{ mol FeCl}_3$

$$\frac{0.27 \text{ M}}{1} \times \frac{1 \text{ L}}{1000 \text{ mL}} \times 518 \text{ mL} = 0.14 \text{ mol FeCl}_3 \quad \boxed{= 0.420 \text{ mol Cl}^-}$$

- 2) What is the molarity of 0.42 mol of NaCl in enough water to make 4.0 L of solution?

- 3) How many moles of H^+ are in 720 mL of 0.41 M nitric acid, HNO_3 ?

- 4) Which of the following ionic compounds are soluble (dissolve) in water?

- a) PbCl_2
- b) FeI_3
- c) NH_4F
- d) $\text{Pb}(\text{NO}_3)_2$

- 5) Which of the following compounds are soluble?

- a) $\text{Al}(\text{OH})_3$
- b) NaOH
- c) KBr
- d) BaSO_4
- e) K_2SO_4

- 6) State whether the salts below are (i) soluble, insoluble;

(ii) electrolyte or nonelectrolyte (charge)

(iii) contains polyatomic ion or not?



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