

Potentially useful data:

Avogadro's constant: $6.022 \times 10^{23} \text{ mol}^{-1}$

1. Determine the set of smallest whole numbers required to balance the chemical equation below:



- | | <i>x</i> | <i>y</i> | <i>z</i> | <i>w</i> |
|----|----------|----------|----------|----------|
| a) | 1 | 2 | 1 | 2 |
| b) | 1 | 3 | 1 | 2 |
| c) | 2 | 3 | 2 | 4 |
| d) | 2 | 5 | 2 | 2 |
| e) | 4 | 6 | 2 | 8 |

2. Which statement about a central atom with four electron groups is FALSE?

- a) The central atom can be a nonmetal.
- b) The electron geometry must be the same as the molecular geometry.
- c) The central atom can be part of a polyatomic ion.
- d) The central atom could be part of a trigonal pyramidal molecular geometry.
- e) The central atom could be part of a bent molecular geometry.

3. Consider the reaction described by the chemical equation below. Estimate the minimum amount (in grams) of TiCl_4 needed to produce 132 millimoles of MgCl_2 .



- a) 6.26 g
- b) 12.5 g
- c) 13.2 g
- d) 25.0 g
- e) 50.1 g

4. Which molecule is nonpolar?

- a) CH_2Cl_2
- b) SO_2
- c) H_2S
- d) SF_4
- e) None. All of the above are polar molecules.

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