

**Part I: Potential Energy Diagrams**

Read pages 527-529 in *Chemistry: the Central Science*.

The energy changes involved in chemical reactions can be conveniently examined using potential energy diagrams. The total potential energy of all the reactants in the reaction is illustrated at the left of the diagram, the total potential energy of the products is at the right side, and the activation energy change or changes are illustrated in between.

On the grids below, plot energy diagrams illustrating the reactions described at left.

**Reaction 1:**

Potential energy of reactants = 250 kJ

Potential energy of activated complex = 350 kJ

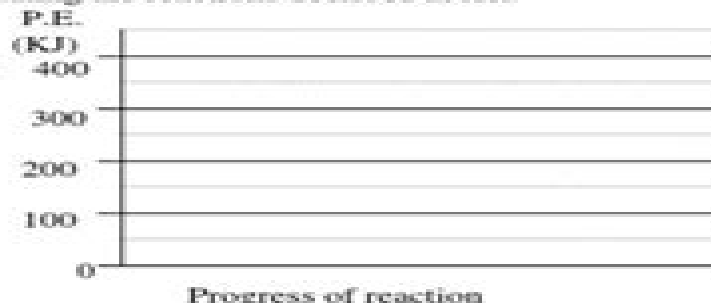
Potential energy of products = 300 kJ

Is this reaction exothermic or endothermic?

How can you tell?

What is the value of  $\Delta H$  for this reaction?

What is the activation energy for this reaction?



In the presence of a catalyst for this reaction, how would the diagram change? Explain which of the following would be changed, and how: energy of reactants, energy of products, energy of activated complex, rate of reaction.

**Reaction 2:**

Potential energy of reactants = 350 kJ

Activation energy = 100 kJ

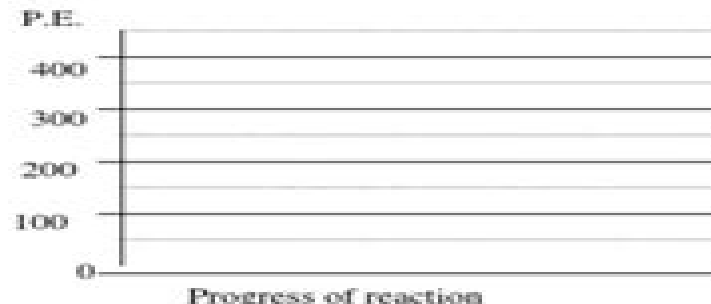
Potential energy of products = 250 kJ

Is this reaction exothermic or endothermic?

How can you tell?

What is the value of  $\Delta H$  for this reaction?

What is the energy of the activated complex for this reaction?



Suppose the concentrations of reactants were increased. Explain which of the following would be changed, and how: energy of reactants, energy of products, energy of activated complex, rate of reaction.

**Selected Answers:** Rxn 1: Endothermic... products at higher PE than reactants.  $\Delta H = +50$  kJ,  $E_a = 100$  kJ. With catalyst,  $E_a$  would be less, increasing the rate of reaction. Rxn 2: exothermic... products at lower PE.  $\Delta H = -100$  kJ. PE of activated complex = 450 kJ. Conc. Changes would affect the rate only. Rxn 4 (next page): PE of reactants = 100 kJ, of activated complex = 350 kJ, of products = 50 kJ.  $\Delta H = -50$  kJ,  $E_a = 250$  kJ

# Chem 12chapter 14 Name

**Ferdinand Gerhard Wiechmann**

A red circular graphic with a gradient, appearing as a stylized arrow or a partial circle, located to the right of the name bar.

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**Elements of Chemistry** William Allen Miller, 1880 *The Etymology of Chemical Names*, 2019-10-08 Etymology of Chemical Names gives an overview of the development of the current chemical nomenclature tracing its sources and changing rules as chemistry progressed over the years This book is devoted to provide a coherent picture how the trivial and systematic names shall be used and how the current IUPAC rules help to reconcile the conflicting demands

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**Inorganic chemistry** William Allen Miller, 1880 *Elements of Chemistry: Organic chemistry* William Allen Miller, 1880

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thanks are due to the students and colleagues who provided feedback and encouragement during its development We sincerely hope that Environmental Chemistry Volume II will serve as a valuable resource for undergraduate and postgraduate students researchers and anyone interested in the chemistry of our environment

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